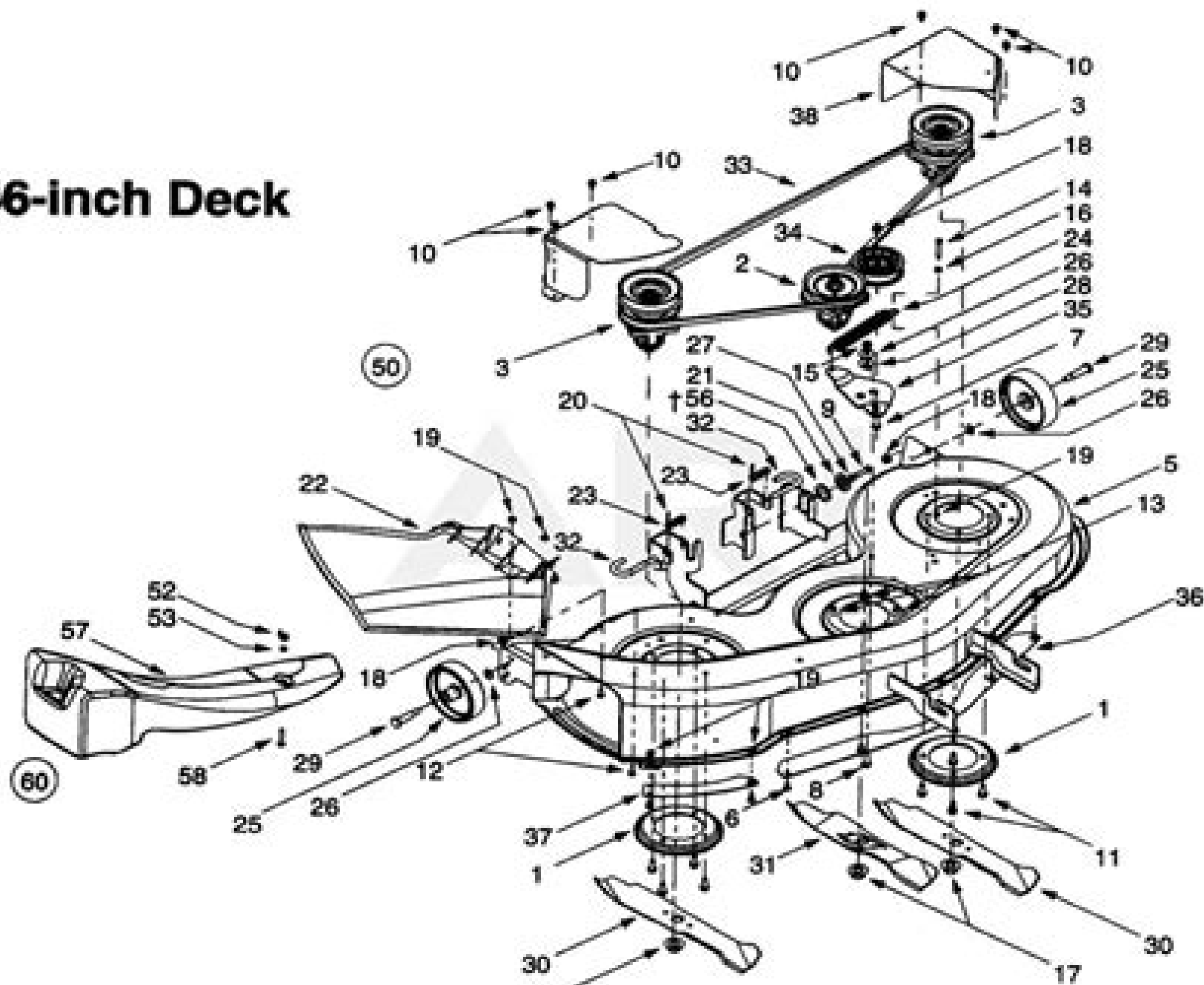


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reaction a redox reaction explain your answer $2\text{NaCl aq} + \text{Pb} \rightarrow \text{Pb} + 2\text{NaCl}$ $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$ $\text{Pb} + 2\text{HCl} \rightarrow \text{PbCl}_2 + 2\text{H}_2$

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web write the oxidation and reduction reactions for each redox reaction the first one is done for you $23 \text{ Fe}^{2+} + \text{Co} \rightarrow \text{Co}^{2+} + \text{Fe}$
oxidation $\text{Co} \rightarrow \text{Co}^{2+} + 2\text{e}^{-}$ reduction $\text{Fe}^{2+} + 2\text{e}^{-} \rightarrow \text{Fe}$ $24 \text{ 3 Ag} + \text{Ni}^{3+} \rightarrow 3 \text{ Ag} + \text{Ni}$ oxidation reduction $25 \text{ Cu}^{2+} + \text{Pb} \rightarrow \text{Pb}^{2+} + \text{Cu}$ oxidation reduction 26 O_2

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 $5\text{Co} + \text{I}_2 \rightarrow 5\text{Co}^{2+} + \text{I}_2$ $5 \text{ Co} + \text{I}_2 \rightarrow 5 \text{ Co}^{2+} + \text{I}_2$

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