

**Table K
Selected Polyatomic Ions**

Formula	Name	Formula	Name
ClO_4^-	perchlorate	ClO_2^-	chlorite
ClO_3^-	chlorate	ClO_2^+	dichlorine oxide
NH_4^+	ammonium	AsO_4^{3-}	arsate
$\text{C}_2\text{H}_3\text{O}_2^-$ CH_3COO^-	acetate	PO_4^{3-}	phosphate
CN^-	cyanide	NO_3^-	nitrate
CO_3^{2-}	carbonate	O_2^{2-}	peroxide
HCO_3^-	hydrogen carbonate	OH^-	hydroxide
$\text{C}_2\text{O}_4^{2-}$	oxalate	PO_3^{2-}	phosphite
CrO_4^{2-}	chromate	NO_2^-	nitrite
CrO_2^-	chromite	NO_2^+	nitron
CrO_3	chromic acid	NO_3^+	nitron
$\text{Cr}_2\text{O}_7^{2-}$	dichromate	HPO_3^{2-}	hydrogen phosphite
Cr_2O_3	chromium(III) oxide	$\text{S}_2\text{O}_3^{2-}$	thiosulfate

**Table L
Solubility Guidelines for Aqueous Solutions**

Basic Most Forms Soluble Compounds	Exceptions	Basic Most Forms Soluble Compounds*	Exceptions
Group 1 ions (Li^+ , Na^+ , etc.)		acetate (CH_3CO_2^-)	when combined with Cu^{2+} or Pb^{2+} ions
ammonium (NH_4^+)		chromate (CrO_4^{2-})	when combined with Group 1 ions, Ca^{2+} , Sr^{2+} , or ammonium (NH_4^+)
nitrate (NO_3^-)		phosphate (PO_4^{3-})	when combined with Group 1 ions or ammonium (NH_4^+)
hydrogen carbonate (HCO_3^-)		oxide (O^{2-})	when combined with Group 1 ions or ammonium (NH_4^+)
halides (F^- , Cl^-)		hydroxide (OH^-)	when combined with Group 1 ions, Ca^{2+} , Sr^{2+} , Ba^{2+} , or ammonium (NH_4^+)
halides (Br^- , I^-)	when combined with Ag^+ , Pb^{2+} , or Hg_2^{2+}		
sulfates (SO_4^{2-})	when combined with Ag^+ , Ca^{2+} , Sr^{2+} , Ba^{2+} , or Pb^{2+}		

*Transition metal ions may have solubility as Hg_2^{2+}

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