

Measurements Using Electrochemical Cells and Electroplating

The basic counting unit in chemistry, the mole, has a special name, *Avogadro's number*, in honor of the Italian scientist Amadeo Avogadro (1776-1856). The commonly accepted definition of Avogadro's number is the number of atoms in exactly 12 g of the isotope ^{12}C , and the quantity itself is $6.02214199(47) \times 10^{23}$.¹

A bit of information about Avogadro seems appropriate. His full name was Lorenzo Romano Amedeo Carlo Avogadro (almost a mole of letters in his name). He was a practicing lawyer until 1806 when he began his new career teaching physics and math at the University of Turin, where he was later promoted to the chair of physical chemistry. In 1811, Avogadro published a paper in the *Journal de Physique*, entitled "Essay on a Manner of Determining the Relative Masses of the Elementary Molecules of Bodies, and the Proportions in Which They Enter into These Compounds," which pretty much says it all. This paper includes the statement that has come to be regarded as Avogadro's Hypothesis:

The first hypothesis to present itself in this connection, and apparently even the only admissible one, is the supposition that the number of integral molecules in any gases is always the same for equal volumes, or always proportional to the volumes. Indeed, if we were to suppose that the number of molecules contained in a given volume were different for different gases, it would scarcely be possible to conceive that the law regulating the distance of molecules could give in all cases relations as simple as those which the facts just detailed compel us to acknowledge between the volumes and the number of molecules.

In this experiment, you will confirm Avogadro's number by conducting an electrochemical process called *electrolysis*. In electrolysis, an external power supply is used to drive an otherwise *nonspontaneous* reaction. You will use a copper strip and a zinc strip as the electrodes, placed in a beaker of sulfuric acid. You will make the cell electrolytic by using the copper strip as the anode and the zinc strip as the cathode. By determining the average current used in the reaction, along with the knowledge that all of the copper ions formed are the 2^+ cations, you will calculate the number of atoms in one mole of copper and compare this value with Avogadro's number.

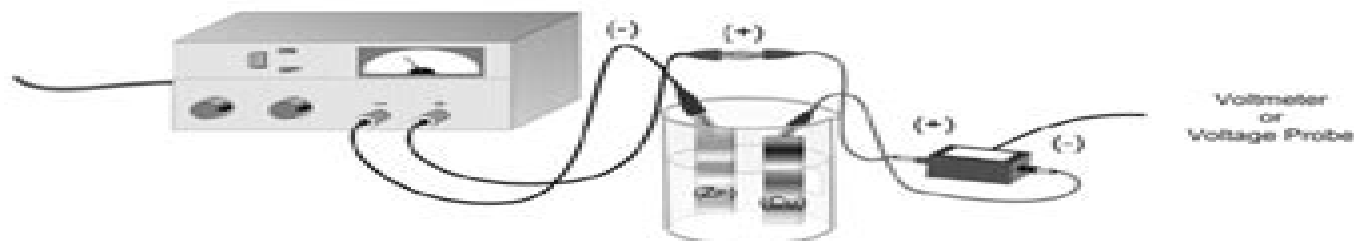


Figure 1

¹ CRC Handbook of Chemistry and Physics, 82nd edition, p. 1-7.

Measurements Using Electrochemical Cells And Electroplating

Kevin R. Reel



Measurements Using Electrochemical Cells And Electroplating:

Embracing the Song of Expression: An Mental Symphony within **Measurements Using Electrochemical Cells And Electroplating**

In a global used by screens and the ceaseless chatter of quick transmission, the melodic splendor and mental symphony developed by the written term often diminish into the background, eclipsed by the constant noise and interruptions that permeate our lives. But, situated within the pages of **Measurements Using Electrochemical Cells And Electroplating** a charming fictional value brimming with fresh feelings, lies an immersive symphony waiting to be embraced. Crafted by an elegant composer of language, that fascinating masterpiece conducts visitors on an emotional trip, skillfully unraveling the hidden tunes and profound influence resonating within each carefully constructed phrase. Within the depths with this touching review, we shall explore the book is main harmonies, analyze its enthralling writing style, and surrender ourselves to the profound resonance that echoes in the depths of readers souls.

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