

C. Melting Point Determination

Recrystallized Benzoic acid	Trials	Melting Point (°C)	Complete Liquidation (°C)	Melting point range (°C)	Change in color	Shrinking of size
	1	122	127	122-127	Simultaneously white to colorless	Gradually shrinks up to 100°C up to melting point
	2	122	127	122-127	Simultaneously white to colorless	Gradually shrinks up to 100°C up to melting point

Benzoic samples	Melting point range (°C)
Recrystallized	122-127
Crude	126-128

- As the crystals were heated, there is a continuous but gradual shrinking of the size of the crystals as evidenced by the diminished or decreasing height of the sample observed in the capillary tubes for both trials accompanied by a change in color as indicated. Also, it was observed that after removal of the capillary tubes after each trial from the hot oil bath, the purified benzoic crystals immediately solidified, that is, returned to the original solid phase.

IV. Discussion and Interpretation of Results

Recrystallization from a solvent depends primarily on the fact that more solids are more soluble in hot solvents than cold solvents, and second, different materials are soluble in different solvents in various solvents. In order for recrystallization to be successful, there should be a large difference in the solubility of the solid material in hot solvents compared to cold solvents. A material can be purified by recrystallization when both the desired substance and the impurity have similar solubilities, but only when the impurity represents a small fraction of the total solid. The desired substance will crystallize on cooling, but the impurity will not.

The most crucial and time-consuming feature of recrystallization is the selection of solvent because important considerations should be observed. In order to choose a good recrystallization solvent, the solubility of the solute should be compared in as many solvents, sparingly soluble only in the solvent at room temperature. Thus, from this relative description, we have to consider the principle "like dissolves like" which refers to the polarity of both the solute and the solvent. Substances of same polarity tend to dissolve one another more at room temperature, which therefore contributes the very first condition of choosing a good recrystallization solvent.

In the experiment, benzoic acid, whose structural formula is given below, was used as the sample.

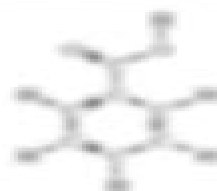


Figure 1. Structural formula of Benzoic acid

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F Rizvi



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Cooperative Chemistry Lab Manual Cooper, 2005-02 The laboratory course described in the lab manual emphasizes experimental design data analysis and problem solving Inherent in the design is the emphasis on communication skills both written and oral Students work in groups on open ended projects in which they are given an initial scenario and then asked to investigate a problem There are no formalized instructions and students must plan and carry out their own investigations

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Table of Contents Organic Chemistry Lab Manual Melting Point Answers

1. Understanding the eBook Organic Chemistry Lab Manual Melting Point Answers
 - The Rise of Digital Reading Organic Chemistry Lab Manual Melting Point Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Organic Chemistry Lab Manual Melting Point Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Organic Chemistry Lab Manual Melting Point Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Organic Chemistry Lab Manual Melting Point Answers
 - Personalized Recommendations
 - Organic Chemistry Lab Manual Melting Point Answers User Reviews and Ratings

- Organic Chemistry Lab Manual Melting Point Answers and Bestseller Lists
- 5. Accessing Organic Chemistry Lab Manual Melting Point Answers Free and Paid eBooks
 - Organic Chemistry Lab Manual Melting Point Answers Public Domain eBooks
 - Organic Chemistry Lab Manual Melting Point Answers eBook Subscription Services
 - Organic Chemistry Lab Manual Melting Point Answers Budget-Friendly Options
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 - Organic Chemistry Lab Manual Melting Point Answers Compatibility with Devices
 - Organic Chemistry Lab Manual Melting Point Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Organic Chemistry Lab Manual Melting Point Answers
 - Highlighting and Note-Taking Organic Chemistry Lab Manual Melting Point Answers
 - Interactive Elements Organic Chemistry Lab Manual Melting Point Answers
- 8. Staying Engaged with Organic Chemistry Lab Manual Melting Point Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Organic Chemistry Lab Manual Melting Point Answers
- 9. Balancing eBooks and Physical Books Organic Chemistry Lab Manual Melting Point Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Organic Chemistry Lab Manual Melting Point Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Organic Chemistry Lab Manual Melting Point Answers
 - Setting Reading Goals Organic Chemistry Lab Manual Melting Point Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Organic Chemistry Lab Manual Melting Point Answers
 - Fact-Checking eBook Content of Organic Chemistry Lab Manual Melting Point Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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