



Perform the following calculations, being sure to give the answer with the correct number of significant digits.

2. Calculate the temperature change that occurs when 364 cal of heat are added to 1.39 kg of ethanol. ($C_{\text{ethanol}} = 2.42 \text{ J/g}^\circ\text{C}$)

$$q = 364 \text{ cal} \times \frac{4.184 \text{ J}}{1 \text{ cal}} = 1522.976 \text{ J}$$

$$m = 1.39 \text{ kg}$$

$$C = 2.42 \text{ J/g}^\circ\text{C}$$

$$\Delta T = ?$$

Organic Chemistry Specific Heat Problems

**Morris Hein, Scott Pattison, Susan
Arena, Leo R. Best**



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