









## Chapter 7 – Alkenes: Structure and Reactivity

## Solutions to Problems

- 7.1 Because two hydrogens must be removed from a saturated compound to introduce an unsaturation, a compound's degree of unsaturation refers to the number of pairs of hydrogens by which its formula differs from that of the corresponding saturated compound. For example, a saturated alkane with four carbons has the formula  $C_4H_{10}$ . The compound in (a),  $C_4H_8$ , which has two fewer (or one pair fewer) hydrogens, may have a double bond or a ring.  $C_8H_{14}$  thus has a degree of unsaturation of 1.

Compound	Degree of Unsaturation		Structures
(a) $C_4H_8$	1	$CH_3CH=CHCH_3$ 	$CH_3CH_2CH=CH_2$ $(CH_3)_2C=CH_2$ 
(b) $C_4H_6$	2	$CH_2=CHCH=CH_2$ $CH_3CH_2C\equiv CH$ 	$CH_3CH-C\equiv CH_2$ $CH_3C\equiv CCCH_3$    
(c) $C_3H_4$	2	$H_2C=C=CH_2$	$CH_3C\equiv CH$ 

- 7.2 Unlike the hydrocarbons in the previous problems, the compounds in this problem contain additional elements. Review the rules for these elements.
- Subtract one hydrogen for each nitrogen present to find the formula of the equivalent hydrocarbon –  $C_6H_4$ . Compared to the alkane  $C_6H_{14}$ , the compound of formula  $C_6H_4$  has 10 fewer hydrogens, or 5 fewer hydrogen pairs, and has a degree of unsaturation of 5.
  - $C_6H_5NO_2$  also has 5 degrees of unsaturation because oxygen does not affect the equivalent hydrocarbon formula of a compound.
  - A halogen atom is equivalent to a hydrogen atom in calculating the equivalent hydrocarbon formula. For  $C_8H_5Cl_3$ , the equivalent hydrocarbon formula is  $C_8H_{12}$ , and the degree of unsaturation is 3.
  - $C_9H_{16}Br_2$  – one degree of unsaturation.
  - $C_{10}H_{12}N_2O_3$  – 6 degrees of unsaturation.
  - $C_{20}H_{30}ClN$  – 5 degrees of unsaturation.

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