

## TYPES OF ORGANIC COMPOUNDS

HYDROCARBON	O-ADDED	>C=O ADDED	>COO ADDED	NITROGEN ADDED	SULFUR ADDED
<b>ALKANE</b> $\text{C}-\text{C}$ <ul style="list-style-type: none"> <li>ethane: <math>\text{C}_2\text{H}_6</math></li> <li>methyl (Me): <math>-\text{CH}_3</math></li> <li>ethyl (Et): <math>-\text{C}_2\text{H}_5</math></li> </ul> <b>ALKENE</b> $\text{C}=\text{C}$ <ul style="list-style-type: none"> <li>ethene: <math>\text{C}_2\text{H}_4</math></li> <li>diene: two <math>\text{C}=\text{C}</math></li> <li>triene: three <math>\text{C}=\text{C}</math></li> </ul> <b>ALKYNE</b> $\text{C}\equiv\text{C}$ <ul style="list-style-type: none"> <li>ethyne: <math>\text{C}_2\text{H}_2</math></li> </ul> <b>AROMATIC</b> <ul style="list-style-type: none"> <li>benzene: <math>\text{C}_6\text{H}_6</math></li> <li>aromatic: <math>\text{C}_6\text{H}_5</math> (Ar-)</li> </ul>	<b>ALCOHOL</b> $\text{C}-\text{OH}$ <ul style="list-style-type: none"> <li>methanol: <math>\text{Me}-\text{OH}</math> (methyl alcohol)</li> <li>phenol: <math>\text{Ar}-\text{OH}</math></li> <li>diol: glycol: (2 -OH)</li> <li>glycerol: (3 -OH)</li> </ul> <b>ETHER</b> $\text{C}-\text{O}-\text{C}$ <ul style="list-style-type: none"> <li>ethoxyethane: <math>\text{Et}-\text{O}-\text{Et}</math> (diethyl ether)</li> </ul> <b>EPOXY</b> <ul style="list-style-type: none"> <li>cyclic ether</li> </ul> <b>PEROXIDE</b> $\text{R}-\text{O}-\text{O}-\text{R}'$	<b>ALDEHYDE</b> <ul style="list-style-type: none"> <li>methanal: <math>\text{H}_2\text{CO}</math> (formaldehyde)</li> <li>hexanaldehyde: <math>\text{Ar}-\text{CHO}</math></li> </ul> <b>KETONE</b> <ul style="list-style-type: none"> <li>2-propanone: <math>\text{Me}-\text{CO}-\text{Me}</math> (dimethyl ketone, acetone)</li> <li>diketone: <math>\text{R}-\text{CO}-\text{R}'-\text{CO}-\text{R}''</math></li> </ul>	<b>CARBOXYLIC ACID</b> <ul style="list-style-type: none"> <li>ethanoic acid: <math>\text{Me}-\text{COOH}</math> (acetic acid)</li> <li>acetic acid: <math>\text{Me}-\text{COOH}</math></li> <li>benzoic acid: <math>\text{Ar}-\text{COOH}</math></li> </ul> <b>Dicarboxylic acid</b> $\text{HOOC}-\text{R}-\text{COOH}$ <b>ESTER</b> <ul style="list-style-type: none"> <li>ethyl acetate: <math>\text{Me}-\text{CO}-\text{OEt}</math></li> </ul> <b>Other derivatives:</b> <ul style="list-style-type: none"> <li>Formic acid: <math>\text{H}-\text{CO}-\text{OH}</math></li> <li>Acid anhydride: <math>\text{RCO}-\text{O}-\text{CO}-\text{R}'</math></li> </ul>	<b>AMINE</b> <ul style="list-style-type: none"> <li>methyl amine: <math>\text{H}_3\text{C}-\text{NH}_2</math></li> <li>phenyl amine: <math>\text{Ar}-\text{NH}_2</math> (aniline)</li> <li><math>\text{R}-\text{NH}_2</math> (1°), <math>\text{RR}'\text{NH}</math> (2°), <math>\text{RRR}'\text{N}</math> (3°)</li> </ul> <b>NITRO</b> $\text{R}-\text{NO}_2$ <b>DIAZO</b> $\text{R}-\text{N}=\text{N}$ <b>NITRILE</b> $\text{R}-\text{C}\equiv\text{N}$ <ul style="list-style-type: none"> <li>methane nitrile: <math>\text{Me}-\text{CN}</math></li> </ul> <b>AMIDE</b> <ul style="list-style-type: none"> <li>acetamide: <math>\text{Me}-\text{CO}-\text{NH}_2</math></li> </ul>	<ul style="list-style-type: none"> <li>thiol: <math>\text{R}-\text{SH}</math></li> <li>thioether: <math>\text{R}-\text{S}-\text{R}'</math></li> <li>disulfide: <math>\text{R}-\text{S}-\text{S}-\text{R}'</math></li> <li>thio ester: <math>\text{R}-\text{CO}-\text{SR}'</math></li> <li>sulfonide: <math>\text{R}-\text{SO}-\text{R}'</math></li> <li>sulfone: <math>\text{R}-\text{SO}_2-\text{R}'</math></li> <li>sulfonic acid: <math>\text{R}-\text{SO}_3\text{H}</math></li> </ul> <b>HALOGEN ADDED</b> <ul style="list-style-type: none"> <li>haloalkane: <math>\text{R}-\text{X}</math> (Me-Cl chloromethane)</li> <li>halobenzene: <math>\text{Ar}-\text{X}</math> (chlorobenzene: <math>\text{Ar}-\text{Cl}</math>)</li> <li>acyl halide: <math>\text{R}-\text{CO}-\text{X}</math></li> <li>aryl halide: <math>\text{Ar}-\text{X}</math></li> </ul>

## FORMULAS AND ISOMERS

**Molecular formula:** Elemental symbols with subscripts denote the composition of a compound

**Empirical formula:** Subscripts denote the relative elemental composition

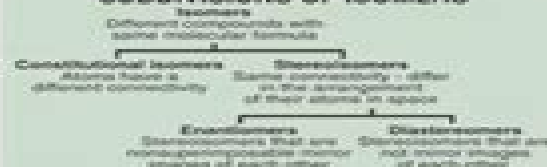
**Graphical depiction:**

- Dash formula:** Diagram all atoms, bonds as dashes
- Bond line formula:** Hide H, show carbon skeleton as lines, other atoms explicit



- Newman Projection:** 2-d depiction
- 3-dimensional:** Wedges of sawhorse denote structure

### SUBDIVISIONS OF ISOMERS



**constitutional isomers:** different bonding connectivity (ex. rings, bonds, branching, substituent positions)

**tautomers:** easily interconverted structural isomers (ex. keto-enol for ketones)

**chiral:** not identical with mirror image

**achiral:** has plane of symmetry (superimposable on mirror-image)

**epimers:** a pair of diastereomers which differ only in the configuration of one atom

**More than 1 chiral center:**

- 1 chiral centers,  $\leq 2^n$  stereoisomers
- more: two chiral centers, 4 isomers; 3 stereoisomers, 1 achiral (mirror-plane)

## FORMULAS AND ISOMERS



**R/S notation:** The four different atoms or groups attached to a central atom are ranked a,b,c,d, by molar mass. The lowest (d) is directed away from the viewer and the sequence of a-b-c produces clockwise (R) or counter-clockwise (S) configuration.

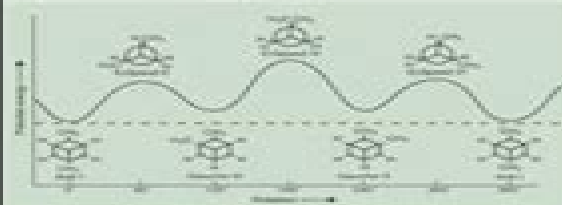
- chiral (optically active):** + or - rotation of plane polarized light, R/S: Opposite effects
- racemic:** 50:50 mixture of stereoisomers (no net optical activity)
- nomenclature:** Note R/S and +/- in the compound name, example: R (+) bromochloromethanol.

**Fischer projection:** Diagram depicts chiral 3-D structure

- molecular conformations:** molecule exhibits structural variation due to free rotation about C-C single bond:



**Newman diagram:** Depict rotation about a C-C bond; eclipsed (high energy), anti (low energy), gauche (intermediate energy).



## COMMON TERMS

- aliphatic:** non-aromatic
- aromatic:** benzene ring
- heterocyclic:** non-carbon atom in the ring structure
- hydrocarbon:** compound of H and C
- paraffin:** alkane
- olefin:** alkene
- saturated:** maximum # of H's (all C-C single bonds)
- unsaturated:** at least one C-C multiple bond

## NOMENCLATURE

**IUPAC:** standard guidelines for naming compounds.

**Nomenclature Strategy:** find longest carbon chain, identify and note location of functional groups and substituents by chain position number.

**Classes of compounds** are defined by the functional group. There are many common names and functional group names. Multiple names are possible.

### CARBON CHAIN PREFIXES

Lat/Cs	Ends	Begin
1	methyl	methyl
2	ethyl	ethyl
3	propyl	propyl
4	butyl	butyl
5	pentyl	pentyl
6	hexyl	hexyl
7	heptyl	heptyl
8	octyl	octyl
9	nonyl	nonyl
10	decyl	decyl

**cyclo-:** ring structure; example: cyclopropane 3-carbon ring molecule

**iso-:** two methyl groups on the

**terminus of a chain**

**n-:** normal straight chain

**i-:** tertiary alkyl group

**vic (vicinally):** two substituents

**on adjacent carbons**

**gem (geminal):** two substituents

**on the same carbon**

**alkene isomers:** cis or trans

**benzene substitution positions:**

ortho (1,2), meta (1,3), para (1,4)

### Carbon Position



# Organic Chemistry Fundamentals Quickstudy Academic

**Clemens Wendtner**



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