

# Reactions of Amines and Epoxides

## Preparation of Amines

Amination of alcohols: Polyalkylation is a serious problem for tertiary hydroxyl groups



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



## Bond Types, Strength, Length, Polarity and Dipole Moment

### Two different bond types

There are two types of bonding: ionic bonding and covalent bonding

- **Ionic bonds** are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron).



- **Covalent bonds** are formed by sharing of electrons between two atoms.



### The Relative Electronegativity of an



### Bonded and Unbonded electron pairs



### Electron movement



### There are two main categories

Definition of electron pair



## Sigma and pi bonds

Covalent bonds are classified in sigma and pi bonds

- **Sigma bonds** are formed by head-on (frontal) overlap of orbitals.
- **pi bonds** are formed by side-on overlap of orbitals.
- **One single bond** is a sigma bond.
- **Two multiple bonds** consist of one sigma and pi bonds.

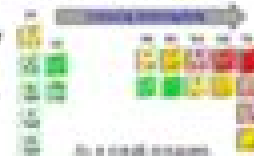


pi bonds are stronger than sigma bonds

## Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity

values, covalent bonds can be polar and nonpolar.



- If the difference in electronegativity is less than 0.5, the electrons are shared equally (nonpolar bond).
- If the difference in electronegativity is between 0.5 and 1.7, the electrons are shared unequally (polar bond).
- If the difference in electronegativity is greater than 1.7, the electrons are not shared (ionic bond).

- A difference of 1.7 or higher can be high, but the difference can be lower (shared) and a lower bond is formed.

## Dipole Moment

The polarity of a bond is described by the dipole moment ( $\mu$ ).



- The dipole moment is defined as a vector pointing from the less electronegative to the more electronegative atom.
- The more electronegative atom has a partial negative charge ( $\delta^-$ ) and the less electronegative atom has a partial positive charge ( $\delta^+$ ).
- $\delta^-$  (delta minus) indicates an electron deficient atom.
- $\delta^+$  (delta plus) indicates an electron rich atom.

## Molecular Dipole

To determine the molecular polarity, the vector sum of the dipoles is used.

The geometry of the molecule is important. Symmetrical molecules are nonpolar, even if the dipole moments are present. A good example is carbon dioxide.



Consider the vector sum of the dipole moments for the entire molecule.

## Bond Length and Bond Strength

The longer the bond, the weaker it is.

Bond	H-H	H-Cl	H-Br	H-I
Length	109 pm	127 pm	141 pm	154 pm
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

Double bonds are stronger than single bonds, and triple bonds are stronger than double bonds.

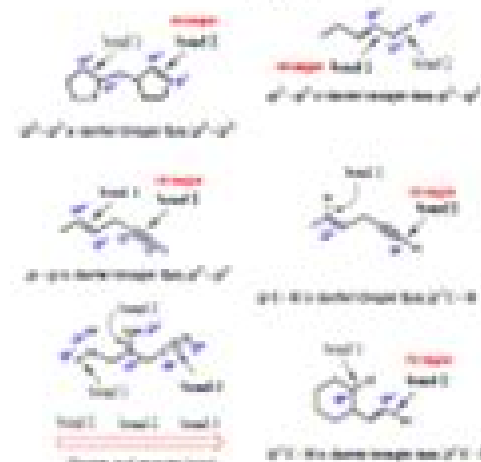


The longer the bond, the weaker it is.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	$sp^3-sp^3$	$sp^2-sp^2$	$sp-sp$
Hybridization	$sp^3-sp^3$	$sp^2-sp^2$	$sp-sp$
Length	154 pm	134 pm	121 pm
Strength	376 kJ/mol	472 kJ/mol	591 kJ/mol
Hybridization	$sp^3-sp^2$	$sp^2-sp$	$sp-sp$
C-H Bond	109 pm	108 pm	106 pm
Length	154 pm	134 pm	121 pm
Strength	376 kJ/mol	472 kJ/mol	591 kJ/mol

Higher s character = stronger and shorter bond

## Examples



# Organic Chemistry Study Guide

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