

Topic	Grade	What do you need to know	 or 	What do You have to work on?
Atomic Structure	A	<ul style="list-style-type: none"> Why an atom is neutral in terms of sub-atomic particles Outline isotopes from data about numbers of protons and neutrons 		
	B	<ul style="list-style-type: none"> Know the charges and masses of protons, electrons and neutrons. Know how isotopes are the same but have different mass number. 		
	C	<ul style="list-style-type: none"> Describe the arrangement of an atoms electron structure - 2.8.8.2 Describe the arrangement of elements in the periodic table. 		
Ionic Bonding	A	<ul style="list-style-type: none"> Explain ionic bonding using the 'dot and cross' model. Explain the melting point of NaCl by the structure of and bonding. Explain why the melting point of NaCl is lower than that of MgO. 		
	B	<ul style="list-style-type: none"> Predict the properties of substances with giant ionic structures. Understand that atoms with a full outer shell are electronically stable. 		
	C	<ul style="list-style-type: none"> Ionic bonds are formed from metals and non-metals Electrons are donated and received. Metals always donate the electron(s). Non- metals always receive the electron(s) 		
The Periodic Table	A	<ul style="list-style-type: none"> Explain some covalent bonds using a 'dot and cross' model Explain double covalent binds using a 'dot and cross' model 		
	B	<ul style="list-style-type: none"> Know why CO₂ and H₂O have low melting points 		
	C	<ul style="list-style-type: none"> Covalent bonds always have non-metals and share electrons Know that the group number is the same as the number of electrons in the outer shell. Be able to find out how many shells an atom has from the period it is in. 		
The Group 1 Elements	A	<ul style="list-style-type: none"> Construct a balanced symbol equation for Cu²⁺, Fe²⁺ and Fe³⁺ to make Cu(OH)₂, Fe(OH)₂, Fe(OH)₃. Reactivity with water gives NaOH and KOH. Know the process is oxidation. 		
	B	<ul style="list-style-type: none"> Why group 1 elements form similar positive ions. Construct a balance symbol equation of group 1 with water 		
	C	<ul style="list-style-type: none"> Metal + Water → Metal Hydroxide + Hydrogen Know how to use a flame test to identify potassium, lithium and sodium compounds. 		
The Group 7 Elements	A	<ul style="list-style-type: none"> Predict if displacement reactions will occur with other halogens Explain the trend of reactivity of group 7 by ideas of electron gain Explain why a process is reduced from its ionic 		

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Table of Contents Ocr Gateway Separate Sciences B C4 Checklist

1. Understanding the eBook Ocr Gateway Separate Sciences B C4 Checklist
 - The Rise of Digital Reading Ocr Gateway Separate Sciences B C4 Checklist
 - Advantages of eBooks Over Traditional Books
2. Identifying Ocr Gateway Separate Sciences B C4 Checklist
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ocr Gateway Separate Sciences B C4 Checklist
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ocr Gateway Separate Sciences B C4 Checklist
 - Personalized Recommendations
 - Ocr Gateway Separate Sciences B C4 Checklist User Reviews and Ratings
 - Ocr Gateway Separate Sciences B C4 Checklist and Bestseller Lists
5. Accessing Ocr Gateway Separate Sciences B C4 Checklist Free and Paid eBooks

- Ocr Gateway Separate Sciences B C4 Checklist Public Domain eBooks
- Ocr Gateway Separate Sciences B C4 Checklist eBook Subscription Services
- Ocr Gateway Separate Sciences B C4 Checklist Budget-Friendly Options
- 6. Navigating Ocr Gateway Separate Sciences B C4 Checklist eBook Formats
 - ePub, PDF, MOBI, and More
 - Ocr Gateway Separate Sciences B C4 Checklist Compatibility with Devices
 - Ocr Gateway Separate Sciences B C4 Checklist Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ocr Gateway Separate Sciences B C4 Checklist
 - Highlighting and Note-Taking Ocr Gateway Separate Sciences B C4 Checklist
 - Interactive Elements Ocr Gateway Separate Sciences B C4 Checklist
- 8. Staying Engaged with Ocr Gateway Separate Sciences B C4 Checklist
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ocr Gateway Separate Sciences B C4 Checklist
- 9. Balancing eBooks and Physical Books Ocr Gateway Separate Sciences B C4 Checklist
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ocr Gateway Separate Sciences B C4 Checklist
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ocr Gateway Separate Sciences B C4 Checklist
 - Setting Reading Goals Ocr Gateway Separate Sciences B C4 Checklist
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ocr Gateway Separate Sciences B C4 Checklist
 - Fact-Checking eBook Content of Ocr Gateway Separate Sciences B C4 Checklist
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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