

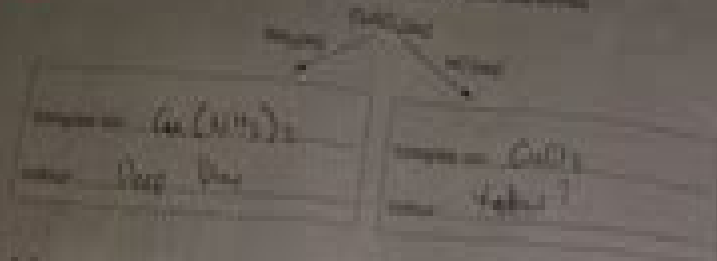
Question 1

Answer all the questions

The reaction is about the synthesis of copper complexes and ligands.

(a) The reaction shows the synthesis of a copper complex.

In the boxes, write the formula and name of the complex and ligand.



(b) Cu^{2+} ions form a complex ion B with two ethylenediamine ions and two water molecules. The ethylenediamine ion is a bidentate ligand.

The skeletal formula of the ethylenediamine ion is shown in Fig. 1.1 below.



Fig. 1.1

(c) What is meant by the term bidentate ligand?

A ligand with two donor atoms that can form two coordinate bonds with a central metal ion.

(d) The complex ion A is shown in Fig. 1.2 below.

The structure of one of the ethylenediamine ligands is shown below. The charge has been given.



Complex A

Complete the 3D diagram of the other two geometries of A. You do not need to include any charges.

Indicate with ticks whether the stereoisomers are cis, trans, optical or a combination of these types.

Diagram														
Type	<table border="1"> <tr> <td>cis</td> <td></td> </tr> <tr> <td>trans</td> <td>✓</td> </tr> <tr> <td>optical</td> <td></td> </tr> </table>	cis		trans	✓	optical		<table border="1"> <tr> <td>cis</td> <td>✓</td> </tr> <tr> <td>trans</td> <td></td> </tr> <tr> <td>optical</td> <td>✓</td> </tr> </table>	cis	✓	trans		optical	✓
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(e) What is the empirical formula, including the charge, of the complex ion A?

$[Cu(NH_3)_2(en)_2]^{2+}$

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Janet Renshaw, Ted Lister



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