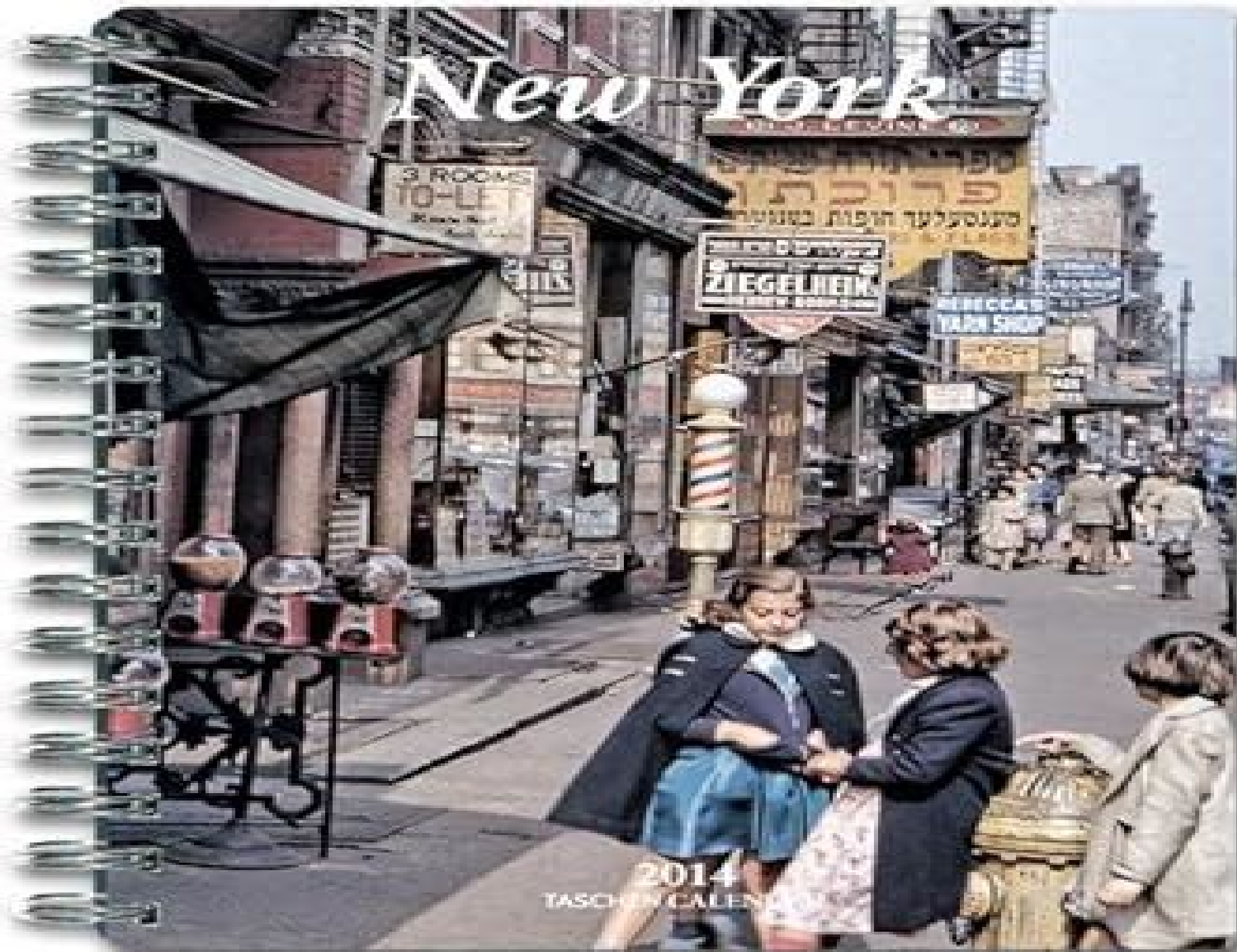


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New York Diary 2014 Taschen, 2013-08 TASCHEN Diary New York 2014 About the Series TASCHEN's Spiral Diaries are beautiful to look at and a joy to use Every week brings a new picture and a fresh page to fill with your life's events Useful features include national holidays year planners and a lay flat spiral binding **Hunter S. Thompson** Kevin T.

McEneaney, 2016-06-09 A decade after Hunter S Thompson's death his books including *Hell's Angels* *The Curse of Lono* *The Great Shark Hunt* and *Rum Diary* continue to sell thousands of copies each year and previously unpublished manuscripts of his still surface for publication While Thompson never claimed to be a great writer he did invent a new literary style gonzo that has been widely influential on both literature and journalism Though Thompson and his work engendered a significant even rabid following relatively little analysis has been published about his writing In *Hunter S Thompson Fear Loathing and the Birth of Gonzo* Kevin T McEneaney examines the intellectual background of this American original providing biographical details and placing Thompson within a larger social and historical context A significant portion of this book is devoted to the creation reception and legacy of his most important works particularly *Fear and Loathing in Las Vegas* In addition to discussing influences on Thompson's work including Homer Nietzsche Spengler Melville Twain Hemingway Fitzgerald Kerouac and others as well as the writers Thompson influenced McEneaney also explains the literary origins of gonzo With new biographical information about Thompson and an examination of his writing techniques this book provides readers with a better understanding of the journalist and novelist A look beyond the larger than life public persona Hunter S Thompson *Fear Loathing and the Birth of Gonzo* will be of great interest to fans of Thompson's work as well as to those wanting to know more about gonzo journalism and literature *Forthcoming Books* Rose Arny, 2000 **Diary of a Woman from New York**, 1886 **My Private Journal** Metropolitan Museum of Art (New York, N.Y.), 1983

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for the nitration 4 01 ml of nitric acid 4 00 ml of sulfuric acid and 4 56 ml bromobenzene were used to complete the reaction with these measurements bromobenzene was found to be the limiting reagent since only 4 56 ml of bromobenzene was used the amount of product predicted to be recovered was 8 7533 g

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the aim of the experiment is to determine and study electrophilic aromatic substitution reaction eas particularly nitration of methyl benzoate go give methyl m nitrobenzoate thus determine the weight melting point and percentage yield of the pure methyl m nitrobenzoate abstract

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in this experiment we will carry out the nitration of bromobenzene bromine substituted on an aromatic ring is an ortho para

directing group it will direct an incoming electrophile to the ortho and para positions it is also deactivating meaning that its reactions will be slower than the identical reactions of benzene

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section 20711 lab experiment nitration of bromobenzene introduction the chemical reaction of nitration consists of a nitro group being added to or substituted in a molecule nitration can basically be carried out by a mixture of concentrated nitric acid and sulfuric acid this mixture is useful to obtain the active nitronium ion

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1 go to the restroom and wash the affected area with lots of soap and water the soap in the lab is not suitable for this purpose 2 return to the lab and apply mineral oil to the affected area 3 the summary to this warning is not to touch any part of your body while performing this experiment if you must leave the lab for any reason 1

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in the nitration of bromobenzene a mixture of sulfuric acid and nitric acid react with one another nitric acid the weaker of the two acids will serve as base to form nitronium ion bromobenzene will attack the nitronium ion by nucleophilic attack

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then actual is your actual findings yield actual theoretical $\times 100$ what are the possible side reactions of nitration of bromobenzene dinitration can occur at ortho para or at both ortho positions occurs in overheating meta product happens very rarely because bromine is an ortho para director

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the soap in the lab is not suitable for this purpose 2 return to the lab and apply mineral oil to the affected area theory and explanation of the nitration of bromobenzene if the temperature of the reaction mixture exceeds 60 o

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the bromine substituent in bromobenzene is an ortho para director the para position is favored due to steric effects in the reaction that takes place in this lab a nitro group is substituted onto bromobenzene at the para position to form 1 bromo 4 nitrobenzene

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study with quizlet and memorize flashcards containing terms like reacting sulfuric and nitric acid together produces the nitronium ion alone nitric acid is not a strong enough electrophile to react with bromobenzene sulfuric acid is added to the nitric acid the reverse rxn can overheat and boil easily includes a claisen adaptor and

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over 5 minutes slowly pipet this bromobenzene into the icy cold nitration mixture hno 3 h 2 so 4 in the 125 ml flask and stir continue to stir the reaction mixture in the ice bath for 30 minutes do not heat in a new 250 ml erlenmeyer

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on the right the arenium cation has a nitro set with incredibly small 0.6 angstrom bond distances this will result in an optimization failure as the atoms are even close for the programme to accomodate their electrons demonstration of creating the p arenium cation by to nitration of bromobenzene

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the purpose of this lab was to observe where substituents are added on a benzene ring by way of electrophilic aromatic substitution the electrophilic source in the nitration of bromobenzene comes from a nitronium ion no2 which acts as the electrophilic species

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nitration of bromobenzene is an electrophilic aromatic substitution reaction the electrophile will be formed from 4.0 ml of nitric acid and 4.0 ml of sulfuric acid and will then react with 4.5 ml of bromobenzene this will produce o bromonitrobenzene and p bromonitrobenzene

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