

**Table K  
Selected Polyatomic Ions**

Formula	Name	Formula	Name
$\text{ClO}_4^-$	perchlorate	$\text{ClO}_2^-$	chlorite
$\text{ClO}_3^-$	chlorate	$\text{ClO}_2^+$	dichlorine oxide
$\text{NH}_4^+$	ammonium	$\text{MnO}_4^-$	permanganate
$\text{C}_2\text{H}_3\text{O}_2^-$ $\text{CH}_3\text{COO}^-$	acetate	$\text{NO}_2^-$	nitrite
$\text{CN}^-$	cyanide	$\text{NO}_3^-$	nitrate
$\text{CO}_3^{2-}$	carbonate	$\text{O}_2^{2-}$	peroxide
$\text{HCO}_3^-$	bicarbonate hydrogencarbonate	$\text{OH}^-$	hydroxide
$\text{C}_2\text{O}_4^{2-}$	oxalate	$\text{PO}_4^{3-}$	phosphate
$\text{CrO}_4^{2-}$	chromate	$\text{HCO}_3^-$	hydrogencarbonate
$\text{CrO}_2^-$	chromite	$\text{SO}_3^{2-}$	sulfite
$\text{CrO}_4^{2-}$	chromate	$\text{SO}_3^-$	sulfite
$\text{Cr}_2\text{O}_7^{2-}$	dichromate	$\text{HPO}_4^{2-}$	hydrogencarbonate
$\text{Cr}_2\text{O}_7^{2-}$	dichromate	$\text{S}_2\text{O}_3^{2-}$	thiosulfate

**Table L  
Solubility Guidelines for Aqueous Solutions**

Basic Most Forms Soluble Compounds	Exceptions	Basic Most Forms Soluble Compounds*	Exceptions
Group 1 ions ( $\text{Li}^+$ , $\text{Na}^+$ , etc.)		acetate ( $\text{CH}_3\text{CO}_2^-$ )	when combined with $\text{Cu}^{2+}$ or $\text{Pb}^{2+}$
ammonium ( $\text{NH}_4^+$ )		chromate ( $\text{CrO}_4^{2-}$ )	when combined with Group 1 ions, $\text{Ca}^{2+}$ , $\text{Sr}^{2+}$ , or ammonium ( $\text{NH}_4^+$ )
nitrate ( $\text{NO}_3^-$ )		phosphate ( $\text{PO}_4^{3-}$ )	when combined with Group 1 ions or ammonium ( $\text{NH}_4^+$ )
hydrogencarbonate ( $\text{HCO}_3^-$ )		oxide ( $\text{O}^{2-}$ )	when combined with Group 1 ions or ammonium ( $\text{NH}_4^+$ )
halides ( $\text{F}^-$ , $\text{Cl}^-$ )		hydroxide ( $\text{OH}^-$ )	when combined with Group 1 ions, $\text{Ca}^{2+}$ , $\text{Sr}^{2+}$ , $\text{Ba}^{2+}$ , or ammonium ( $\text{NH}_4^+$ )
halides ( $\text{Br}^-$ , $\text{I}^-$ )	when combined with $\text{Ag}^+$ , $\text{Pb}^{2+}$ , or $\text{Hg}_2^{2+}$		
sulfates ( $\text{SO}_4^{2-}$ )	when combined with $\text{Ag}^+$ , $\text{Ca}^{2+}$ , $\text{Sr}^{2+}$ , $\text{Ba}^{2+}$ , or $\text{Pb}^{2+}$		

\*Ammonium has varying solubility in  $\text{H}_2\text{O}$

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