

**Table K**  
**Selected Polyatomic Ions**

Formula	Name	Formula	Name
$\text{ClO}_4^-$	perchlorate	$\text{ClO}_2^-$	chlorite
$\text{ClO}_3^-$	chlorate	$\text{ClO}_2^+$	dichlorine oxide
$\text{NH}_4^+$	ammonium	$\text{MnO}_4^-$	permanganate
$\text{C}_2\text{H}_3\text{O}_2^-$ $\text{CH}_3\text{COO}^-$	acetate	$\text{NO}_2^-$	nitrite
$\text{CN}^-$	cyanide	$\text{NO}_3^-$	nitrate
$\text{CO}_3^{2-}$	carbonate	$\text{O}_2^{2-}$	peroxide
$\text{HCO}_3^-$	hydrogen carbonate	$\text{OH}^-$	hydroxide
$\text{C}_2\text{O}_4^{2-}$	oxalate	$\text{PO}_4^{3-}$	phosphate
$\text{CrO}_4^{2-}$	chromate	$\text{SCN}^-$	thiocyanate
$\text{CrO}_2^-$	chromite	$\text{SO}_3^{2-}$	sulfite
$\text{CrO}_4^{2-}$	chromate	$\text{SO}_3^{2-}$	sulfite
$\text{CrO}_4^{2-}$	chromate	$\text{HSO}_4^-$	hydrogen sulfate
$\text{CrO}_4^{2-}$	chromate	$\text{S}_2\text{O}_8^{2-}$	persulfate

**Table L**  
**Solubility Guidelines for Aqueous Solutions**

Basic Fact Form Soluble Compounds	Exceptions	Basic Fact Form Soluble Compounds*	Exceptions
Group 1 ions ( $\text{Li}^+$ , $\text{Na}^+$ , etc.)		acetate ( $\text{CH}_3\text{CO}_2^-$ )	when combined with $\text{Cu}^{2+}$ or $\text{Pb}^{2+}$
ammonium ( $\text{NH}_4^+$ )		chromate ( $\text{CrO}_4^{2-}$ )	when combined with Group 1 ions, $\text{Ca}^{2+}$ , $\text{Sr}^{2+}$ , or ammonium ( $\text{NH}_4^+$ )
nitrate ( $\text{NO}_3^-$ )		phosphate ( $\text{PO}_4^{3-}$ )	when combined with Group 1 ions or ammonium ( $\text{NH}_4^+$ )
hydrogen carbonate ( $\text{HCO}_3^-$ )		oxide ( $\text{O}^{2-}$ )	when combined with Group 1 ions or ammonium ( $\text{NH}_4^+$ )
chloride ( $\text{Cl}^-$ )		hydroxide ( $\text{OH}^-$ )	when combined with Group 1 ions, $\text{Ca}^{2+}$ , $\text{Sr}^{2+}$ , or ammonium ( $\text{NH}_4^+$ )
iodide ( $\text{I}^-$ )			
sulfate ( $\text{SO}_4^{2-}$ )	when combined with $\text{Ag}^+$ , $\text{Pb}^{2+}$ , or $\text{Ba}^{2+}$		

# New York State Ratings Guide Chemistry Regents

**RS Peters**



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## **Table of Contents New York State Ratings Guide Chemistry Regents**

1. Understanding the eBook New York State Ratings Guide Chemistry Regents
  - The Rise of Digital Reading New York State Ratings Guide Chemistry Regents
  - Advantages of eBooks Over Traditional Books
2. Identifying New York State Ratings Guide Chemistry Regents
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in a New York State Ratings Guide Chemistry Regents
  - User-Friendly Interface
4. Exploring eBook Recommendations from New York State Ratings Guide Chemistry Regents
  - Personalized Recommendations
  - New York State Ratings Guide Chemistry Regents User Reviews and Ratings
  - New York State Ratings Guide Chemistry Regents and Bestseller Lists
5. Accessing New York State Ratings Guide Chemistry Regents Free and Paid eBooks
  - New York State Ratings Guide Chemistry Regents Public Domain eBooks
  - New York State Ratings Guide Chemistry Regents eBook Subscription Services
  - New York State Ratings Guide Chemistry Regents Budget-Friendly Options

6. Navigating New York State Ratings Guide Chemistry Regents eBook Formats
  - ePub, PDF, MOBI, and More
  - New York State Ratings Guide Chemistry Regents Compatibility with Devices
  - New York State Ratings Guide Chemistry Regents Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of New York State Ratings Guide Chemistry Regents
  - Highlighting and Note-Taking New York State Ratings Guide Chemistry Regents
  - Interactive Elements New York State Ratings Guide Chemistry Regents
8. Staying Engaged with New York State Ratings Guide Chemistry Regents
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers New York State Ratings Guide Chemistry Regents
9. Balancing eBooks and Physical Books New York State Ratings Guide Chemistry Regents
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection New York State Ratings Guide Chemistry Regents
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine New York State Ratings Guide Chemistry Regents
  - Setting Reading Goals New York State Ratings Guide Chemistry Regents
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of New York State Ratings Guide Chemistry Regents
  - Fact-Checking eBook Content of New York State Ratings Guide Chemistry Regents
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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