

- **Thermodynamics** is the study of energy and its transformations.
 - **First Law of Thermodynamics**: Energy cannot be created or destroyed, only transformed.

$$\Delta U = Q - W$$

 where ΔU is the change in internal energy, Q is the heat added to the system, and W is the work done by the system.



- **Second Law of Thermodynamics**: The entropy of an isolated system never decreases.
 - **Entropy** is a measure of the disorder or randomness of a system.
 - **Third Law of Thermodynamics**: The entropy of a perfect crystal at absolute zero is zero.



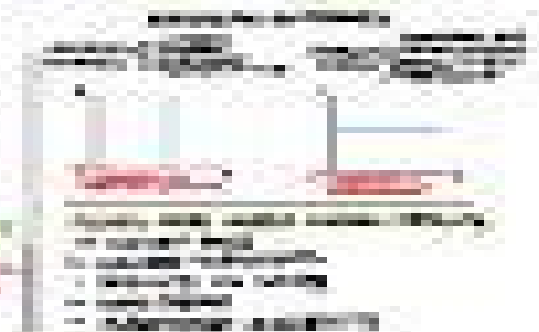
- **Heat Engines** convert heat into work.
 - **Efficiency** is the ratio of work output to heat input.
 - **Carnot Engine** is the most efficient heat engine operating between two temperatures.



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Xiaolong Qi



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