## Unit 3 Energy Changes and Rates of Reaction

#### Solutions to Practice Problems in Chapter 5 Energy Changes

# Calculating the Absorption of Heat (Student textbook page 281)

 How much heat must be added to a 36.2 g sample of iron to increase its temperature by 250.0°C?

#### What Is Required?

You have to calculate the amount of heat that must be added to a sample of iron to increase its temperature by 250.0°C.

#### What Is Given?

You know the mass of iron: m = 36.2 g

You know the change in the temperature:  $\Delta T = 250.0^{\circ}$ C

You have the specific heat capacity of iron from Table 5.1 on page 280 of the student text:  $c_{Fe} = 0.449 \text{ J/g} \cdot ^{\circ}\text{C}$ 

Plan Your Strategy	Act on Your Strategy
Substitute the given information into the formula $Q = mc\Delta T$ . Solve for $Q$ .	$Q = mc\Delta T$ = $(36.2 \text{ f}) \left(0.449 \frac{\text{J}}{\text{f} \cdot 26}\right) (250.026)$ = $4063.45 \text{ J}$ = $4.06 \text{ kJ}$ The iron must gain $4060 \text{ J}$ or $4.06 \text{ kJ}$ of heat energy.

#### Check Your Solution

Check that the information has been substituted correctly. The answer has the correct number of significant digits.

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#### **Table of Contents Mhr Chemistry 12 Solutions**

- 1. Understanding the eBook Mhr Chemistry 12 Solutions
  - The Rise of Digital Reading Mhr Chemistry 12 Solutions
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Mhr Chemistry 12 Solutions
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Mhr Chemistry 12 Solutions
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Mhr Chemistry 12 Solutions
  - Personalized Recommendations
  - $\,\circ\,$  Mhr Chemistry 12 Solutions User Reviews and Ratings
  - Mhr Chemistry 12 Solutions and Bestseller Lists
- 5. Accessing Mhr Chemistry 12 Solutions Free and Paid eBooks

- Mhr Chemistry 12 Solutions Public Domain eBooks
- Mhr Chemistry 12 Solutions eBook Subscription Services
- o Mhr Chemistry 12 Solutions Budget-Friendly Options
- 6. Navigating Mhr Chemistry 12 Solutions eBook Formats
  - o ePub, PDF, MOBI, and More
  - Mhr Chemistry 12 Solutions Compatibility with Devices
  - Mhr Chemistry 12 Solutions Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Mhr Chemistry 12 Solutions
  - Highlighting and Note-Taking Mhr Chemistry 12 Solutions
  - Interactive Elements Mhr Chemistry 12 Solutions
- 8. Staying Engaged with Mhr Chemistry 12 Solutions
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - o Following Authors and Publishers Mhr Chemistry 12 Solutions
- 9. Balancing eBooks and Physical Books Mhr Chemistry 12 Solutions
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Mhr Chemistry 12 Solutions
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Mhr Chemistry 12 Solutions
  - Setting Reading Goals Mhr Chemistry 12 Solutions
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mhr Chemistry 12 Solutions
  - Fact-Checking eBook Content of Mhr Chemistry 12 Solutions
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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