

L'ASPECT ÉNERGÉTIQUE DES TRANSFORMATIONS

QUESTIONS À CHOIX MULTIPLE

1) Quels énoncés sont vrais pour une réaction exothermique ?

- ❶ L'enthalpie des réactifs est moins élevée que celle des produits.
- ❷ La réaction dégage de l'énergie dans son environnement.
- ❸ La variation d'enthalpie est positive.
- ❹ La somme des énergies de liaison est plus élevée dans les produits que dans les réactifs.
- ❺ Le bilan énergétique est du côté des réactifs.
- ❻ Les produits ont nécessairement une enthalpie négative.

- A) ❶, ❷ et ❸ B) ❷, ❸ et ❹ C) ❷ et ❹ D) ❸ et ❹

2) Soit une réaction chimique telle que $E_{a \text{ directe}} > E_{a \text{ inverse}}$. Que peut-on déduire quant à cette réaction ?

- A) $\Delta H_{\text{directe}} > 0$ et $\Delta H_{\text{inverse}} > 0$ B) $\Delta H_{\text{directe}} > 0$ et $\Delta H_{\text{inverse}} < 0$
C) $\Delta H_{\text{directe}} < 0$ et $\Delta H_{\text{inverse}} < 0$ D) $\Delta H_{\text{directe}} < 0$ et $\Delta H_{\text{inverse}} > 0$

3) Quel type de système permet de faire des expériences de calorimétrie valides ?

- A) Un système ouvert B) Un système fermé
C) Un système isolé D) Un système hermétique

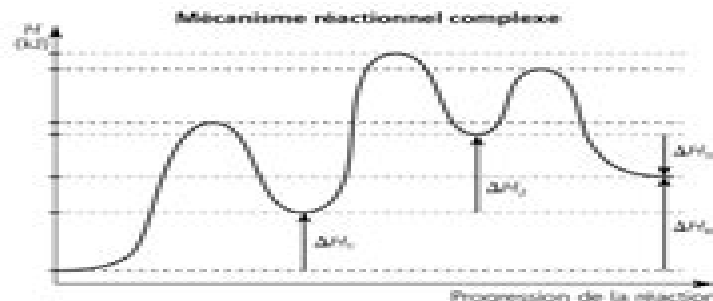
4) Sur le diagramme énergétique d'une réaction exothermique, à quoi correspond le niveau d'enthalpie le plus élevé ?

- A) À l'enthalpie des réactifs B) À l'enthalpie des produits
C) À l'enthalpie du complexe activé D) Ça dépend des situations.

5) Soit le diagramme d'enthalpie ci-contre. D'après la loi de Hess, quelle ou quelles égalités sont vraies ?

- ❶ $\Delta H_{\text{réaction}} = \Delta H_1 + \Delta H_2 + \Delta H_3 + \Delta H_4$
- ❷ $\Delta H_1 = \Delta H_2 + \Delta H_3 + \Delta H_4$
- ❸ $\Delta H_4 = \Delta H_1 - \Delta H_2 - \Delta H_3$
- ❹ $\Delta H_{\text{réaction}} = \Delta H_4$
- ❺ $\Delta H_1 = \Delta H_4 - (\Delta H_2 + \Delta H_3)$
- ❻ $\Delta H_1 + \Delta H_2 + \Delta H_3 + \Delta H_4 = 0$

- A) Toutes B) Aucune
C) ❶ et ❷ D) ❹ et ❺



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Shasha Hu



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