

Thermochemistry Chemistry Questions with Solutions

Q-1: Given that standard molar enthalpies of formation of NO(g) and $\text{NO}_2\text{(g)}$ are respectively 90.3 kJ/mol and 33.2 kJ/mol. Calculate the enthalpy change for the reaction $2\text{NO(g)} + \text{O}_2\text{(g)} \rightarrow 2\text{NO}_2\text{(g)}$.

Answer: $\Delta H = \sum \Delta H_{\text{products}} - \sum \Delta H_{\text{reactants}}$

$$\Delta H = 2(33.2) - [2(90.3) + 0]$$

$$\Delta H = -114.2 \text{ kJ}$$

Note: Enthalpy of formation of free elements like O_2 is always 0.

Q-2: The enthalpy of formation of AgCl is obtained from the enthalpy change from which one of the following process?

- a) $\text{Ag}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{AgCl(s)}$
- b) $\text{Ag(s)} + \frac{1}{2} \text{Cl}_2(\text{g}) \rightarrow \text{AgCl(s)}$
- c) $\text{AgCl(s)} \rightarrow \text{Ag(s)} + \frac{1}{2} \text{Cl}_2(\text{g})$
- d) $\text{Ag(s)} + \text{AuCl} \rightarrow \text{Au(s)} + \text{AgCl(s)}$

Answer: b) $\text{Ag(s)} + \frac{1}{2} \text{Cl}_2(\text{g}) \rightarrow \text{AgCl(s)}$

Explanation: The enthalpy change accompanying the formation of one mole of a pure substance in its standard state from the constituents elements in their standard state is called enthalpy of formation. For Ag, the standard state is solid and for chlorine, it is gas. Thus, the process that gives the enthalpy of formation of AgCl is: $\text{Ag(s)} + \frac{1}{2} \text{Cl}_2(\text{g}) \rightarrow \text{AgCl(s)}$

Q-3: ΔH and ΔE for the reaction, $\text{Fe}_2\text{O}_3\text{(s)} + 3\text{H}_2\text{(g)} \rightarrow 2\text{Fe(s)} + 3\text{H}_2\text{O(l)}$ at constant temperature are related as

- a) $\Delta H = \Delta E$
- b) $\Delta H = \Delta E + RT$
- c) $\Delta H = \Delta E + 3RT$
- d) $\Delta H = \Delta E - 3RT$

Answer: d) $\Delta H = \Delta E - 3RT$

Explanation: For any chemical reaction,

$$\Delta H = \Delta E + \Delta_{\text{ng}}RT$$

Where, Δ_{ng} = total number of moles of gaseous product- total number of moles of gaseous reactants.

Objectives Questions On Thermochemistry

Ranjeet Shahi



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Engineering Chemistry I (for BPUT) Patrizia Bocchetta, 2024 **General Chemistry Workbook** Daniel C. Tofan, 2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions. It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course. It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material.

A Textbook of Physical Chemistry, 6th Edition Sharma K.K. & Sharma L.K., 2016 A Textbook of Physical Chemistry **Internal Combustion Engines** Shyam K. Agrawal, 2006 Salient Features The New Edition Is A Thoroughly Revised Version Of The Earlier Edition And Presents A Detailed Exposition Of The Basic Principles Of Design Operation And Characteristics Of Reciprocating I C Engines And Gas Turbines Chemistry Of Combustion Engine Cooling And Lubrication Requirements Liquid And Gaseous Fuels For Ic Engines Compressors Supercharging And Exhaust Emission Its Standards And Control Thoroughly Explained Jet And Rocket Propulsion Alternate Potential Engines Including Hybrid Electric And Fuel Cell Vehicles Are Discussed In Detail Chapter On Ignition System Includes Electronic Injection Systems For Si And Ci

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Table of Contents Objectives Questions On Thermochemistry

1. Understanding the eBook Objectives Questions On Thermochemistry
 - The Rise of Digital Reading Objectives Questions On Thermochemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Objectives Questions On Thermochemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Objectives Questions On Thermochemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Objectives Questions On Thermochemistry
 - Personalized Recommendations
 - Objectives Questions On Thermochemistry User Reviews and Ratings
 - Objectives Questions On Thermochemistry and Bestseller Lists
5. Accessing Objectives Questions On Thermochemistry Free and Paid eBooks
 - Objectives Questions On Thermochemistry Public Domain eBooks
 - Objectives Questions On Thermochemistry eBook Subscription Services

- Objectives Questions On Thermochemistry Budget-Friendly Options
- 6. Navigating Objectives Questions On Thermochemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Objectives Questions On Thermochemistry Compatibility with Devices
 - Objectives Questions On Thermochemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Objectives Questions On Thermochemistry
 - Highlighting and Note-Taking Objectives Questions On Thermochemistry
 - Interactive Elements Objectives Questions On Thermochemistry
- 8. Staying Engaged with Objectives Questions On Thermochemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Objectives Questions On Thermochemistry
- 9. Balancing eBooks and Physical Books Objectives Questions On Thermochemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Objectives Questions On Thermochemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Objectives Questions On Thermochemistry
 - Setting Reading Goals Objectives Questions On Thermochemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Objectives Questions On Thermochemistry
 - Fact-Checking eBook Content of Objectives Questions On Thermochemistry
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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