AP Chemistry Review Chapter 6 and 7, Atomic Structure and Periodicity

- Be familiar with the electromagnetic spectrum
 - 1. What types of light are shorter in length than microwave? Know the whole order of the spectrum.
 - 2. What type of light helps scientists to learn about electron structure?
 - 3. What creates the emission lines in a spectra for an element? What would cause only certain wavelengths of light to be absorbed by an atom?
- · Convert between frequency and wavelength
 - 4. Calculate the frequency of light emitted by exciting Sr atoms if the wavelength is 461 nm.
 - Calculate the wavelength of light emitted by exciting Na atoms if the frequency is 3.45 X 10¹⁴ s⁻¹
- · Calculate the energy of a wave or an electron
 - Calculate the energy of 1 mole of photons emitted by a Li atom if the wavelength is 670nm.
- Define and understand:
 - Ground state, excited state, Pauli exclusion principle, Heisenberg Uncertainty Principle, Hund's rule.
- Know the maximum number of electrons in and shape of each type of orbital.
 - 8. 3s 5p 4f 6d
- Following the Aufbua principle, determine electron configurations, both longhand and condensed, for elements and ions
- Match each explanation below with the correct letter
 - (A) 1s² 2s² 2p⁵ 3s² 3p⁵
 - (B) 1s2 2s2 2p6 3s2 3p6
 - (C) 1s2 2s2 2p6 2d10 3s2 3p6
 - (D) 1s2 2s2 2p6 3s2 3p6 3d5
 - (E) 1s2 2s2 2p6 3s2 3p6 3d3 4s2
 - 9. An impossible electron configuration
 - 10. The ground-state configuration for the atoms of a transition element
 - 11. The ground-state configuration of a negative ion of a halogen
 - 12. The ground-state configuration of a common ion of an alkaline earth element
 - List 3 species that are isoelectronic
 - 14. Which of the following represents the ground state electron configuration for the Mn³⁺ ion? (Atomic number Mn = 25) (a) 1s² 2s² 2p⁶ 3s² 3p⁶ 3d⁴
 - (b) 1s2 2s2 2p6 3s2 3p6 3d5 4s2 (c) 1s2 2s2 2p6 3s2 3p6 3d2 4s2
 - (d) 1s2 2s2 2p6 3s2 3p6 3d8 4s2 (e) 1s2 2s2 2p6 3s2 3p6 3d3 4s1
- Draw Dot diagrams for elements
 - Boron Calcium Krypton

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