

| Topic                | Grade | What do you need to know   | 😊 or 😞 | What do You have to work on? |
|----------------------|-------|--|--------|------------------------------|
| Atomic Structure     | A     | <ul style="list-style-type: none"> <li>Why an atom is neutral in terms of sub-atomic particles</li> <li>Outline isotopes from data about numbers of protons and neutrons</li> </ul>  |        |                              |
|                      | B     | <ul style="list-style-type: none"> <li>Know the charges and masses of protons, electrons and neutrons.</li> <li>Know how isotopes are the same but have different mass number.</li> </ul>  |        |                              |
|                      | C     | <ul style="list-style-type: none"> <li>Describe the arrangement of an atoms electron structure - 2.8.8.2</li> <li>Describe the arrangement of elements in the periodic table.</li> </ul>   |        |                              |
| Ionic Bonding        | A     | <ul style="list-style-type: none"> <li>Explain ionic bonding using the 'dot and cross' model.</li> <li>Explain the melting point of NaCl by the structure of and bonding.</li> <li>Explain why the melting point of NaCl is lower than that of MgO.</li> </ul>   |        |                              |
|                      | B     | <ul style="list-style-type: none"> <li>Predict the properties of substances with giant ionic structures.</li> <li>Understand that atoms with a full outer shell are electronically stable.</li> </ul>  |        |                              |
|                      | C     | <ul style="list-style-type: none"> <li>Ionic bonds are formed from metals and non-metals</li> <li>Electrons are donated and received.</li> <li>Metals always donate the electron(s).</li> <li>Non- metals always receive the electron(s)</li> </ul>  |        |                              |
| The Periodic Table   | A     | <ul style="list-style-type: none"> <li>Explain some covalent bonds using a 'dot and cross' model</li> <li>Explain double covalent binds using a 'dot and cross' model</li> </ul>   |        |                              |
|                      | B     | <ul style="list-style-type: none"> <li>Know why CO<sub>2</sub> and H<sub>2</sub>O have low melting points</li> </ul>   |        |                              |
|                      | C     | <ul style="list-style-type: none"> <li>Covalent bonds always have non-metals and share electrons</li> <li>Know that the group number is the same as the number of electrons in the outer shell.</li> <li>Be able to find out how many shells an atom has from the period it is in.</li> </ul>                  |        |                              |
| The Group 1 Elements | A     | <ul style="list-style-type: none"> <li>Construct a balanced symbol equation for Cu<sup>2+</sup>, Fe<sup>2+</sup> and Fe<sup>3+</sup> to make Cu(OH)<sub>2</sub>, Fe(OH)<sub>2</sub>, Fe(OH)<sub>3</sub>.</li> <li>Reactivity with water gives NaOH and KOH.</li> <li>Know the process is oxidation.</li> </ul> |        |                              |
|                      | B     | <ul style="list-style-type: none"> <li>Why group 1 elements form similar positive ions.</li> <li>Construct a balance symbol equation of group 1 with water</li> </ul>  |        |                              |
|                      | C     | <ul style="list-style-type: none"> <li>Metal + Water → Metal Hydroxide + Hydrogen</li> <li>Know how to use a flame test to identify potassium, lithium and sodium compounds.</li> </ul>  |        |                              |
| The Group 7 Elements | A     | <ul style="list-style-type: none"> <li>Predict if displacement reactions will occur with other halogens</li> <li>Explain the trend of reactivity of group 7 by ideas of electron gain</li> <li>Explain why a process is reduced from its ionic</li> </ul>  |        |                              |

# Ocr Gateway Separate Sciences B C4 Checklist

**Eliot Attridge, Dorothy Warren, Trevor  
Banham Baker, Neil Dixon, Robert  
Woodcock, Adam Bailey, Steve  
Wiseman**

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## **Table of Contents Ocr Gateway Separate Sciences B C4 Checklist**

1. Understanding the eBook Ocr Gateway Separate Sciences B C4 Checklist
  - The Rise of Digital Reading Ocr Gateway Separate Sciences B C4 Checklist
  - Advantages of eBooks Over Traditional Books
2. Identifying Ocr Gateway Separate Sciences B C4 Checklist
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ocr Gateway Separate Sciences B C4 Checklist
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ocr Gateway Separate Sciences B C4 Checklist
  - Personalized Recommendations
  - Ocr Gateway Separate Sciences B C4 Checklist User Reviews and Ratings
  - Ocr Gateway Separate Sciences B C4 Checklist and Bestseller Lists
5. Accessing Ocr Gateway Separate Sciences B C4 Checklist Free and Paid eBooks
  - Ocr Gateway Separate Sciences B C4 Checklist Public Domain eBooks
  - Ocr Gateway Separate Sciences B C4 Checklist eBook Subscription Services
  - Ocr Gateway Separate Sciences B C4 Checklist Budget-Friendly Options
6. Navigating Ocr Gateway Separate Sciences B C4 Checklist eBook Formats

- ePub, PDF, MOBI, and More
- Ocr Gateway Separate Sciences B C4 Checklist Compatibility with Devices
- Ocr Gateway Separate Sciences B C4 Checklist Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ocr Gateway Separate Sciences B C4 Checklist
  - Highlighting and Note-Taking Ocr Gateway Separate Sciences B C4 Checklist
  - Interactive Elements Ocr Gateway Separate Sciences B C4 Checklist
- 8. Staying Engaged with Ocr Gateway Separate Sciences B C4 Checklist
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ocr Gateway Separate Sciences B C4 Checklist
- 9. Balancing eBooks and Physical Books Ocr Gateway Separate Sciences B C4 Checklist
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ocr Gateway Separate Sciences B C4 Checklist
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Ocr Gateway Separate Sciences B C4 Checklist
  - Setting Reading Goals Ocr Gateway Separate Sciences B C4 Checklist
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ocr Gateway Separate Sciences B C4 Checklist
  - Fact-Checking eBook Content of Ocr Gateway Separate Sciences B C4 Checklist
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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