

Chemistry 12

Solutions Manual Part B

Unit 3 Energy Changes and Rates of Reaction

Solutions to Practice Problems in Chapter 5 Energy Changes

Calculating the Absorption of Heat (Student textbook page 281)

1. How much heat must be added to a 36.2 g sample of iron to increase its temperature by 250.0°C?

What Is Required?

You have to calculate the amount of heat that must be added to a sample of iron to increase its temperature by 250.0°C.

What Is Given?

You know the mass of iron: $m = 36.2 \text{ g}$

You know the change in the temperature: $\Delta T = 250.0^\circ\text{C}$

You have the specific heat capacity of iron from **Table 5.1** on page 280 of the student text: $c_{\text{Fe}} = 0.449 \text{ J/g}^\circ\text{C}$

Plan Your Strategy	Act on Your Strategy
Substitute the given information into the formula $Q = mc\Delta T$. Solve for Q .	$Q = mc\Delta T$ $= (36.2 \text{ g}) \left(0.449 \frac{\text{J}}{\text{g}^\circ\text{C}} \right) (250.0^\circ\text{C})$ $= 4063.45 \text{ J}$ $= 4.06 \text{ kJ}$ The iron must gain 4060 J or 4.06 kJ of heat energy.

Check Your Solution

Check that the information has been substituted correctly. The answer has the correct number of significant digits.

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