

## Assessment and Evaluation

ThoughtLab/ ExpressLab/ Investigation	Curriculum Expectations	Assessment Tools/Techniques	Achievement Chart Category	Learning Skills
ExpressLab: A Penny for Your Isotopes, p. 168	Understanding Basic Concepts <ul style="list-style-type: none"> <li>[OCR 1.02] explain the relationship between isotopic abundance and relative atomic mass</li> </ul>	<ul style="list-style-type: none"> <li>Assessment Checklist 5: Learning Skills (see "Assessment and Evaluation" in the front matter of <i>Teacher's Resource CD-ROM</i>)</li> </ul>	<ul style="list-style-type: none"> <li>Knowledge/Understanding</li> </ul>	<ul style="list-style-type: none"> <li>Teamwork</li> </ul>

## Solutions to Practice Problems

Student Textbook page 170

See *Solutions Manual* for solutions to Practice Problems.

## Section Review Answers

Student Textbook page 170

- The average atomic mass is a weighted average of the atomic masses of the isotopes of potassium. Any given atom would have the atomic mass of one of the isotopes, not the average mass of all the isotopes.

- | Isotope | Atomic Mass (u) | Percentage Abundance |
|---------|-----------------|----------------------|
| Mg-24   | 23.985          | 78.70                |
| Mg-25   | 24.985          | 10.13                |
| Mg-26   | 25.983          | 11.17                |

$$\text{average atomic mass} = \frac{[78.70(23.985) + 10.13(24.985) + 11.17(25.983)] \text{ u}}{100}$$

$$\text{average atomic mass} = \frac{[1887.6 + 253.1 + 290.2] \text{ u}}{100}$$

$$\text{average atomic mass} = 24.31 \text{ u}$$

Therefore, the average atomic mass of magnesium is 24.31 u.

- It is not possible to solve the problem. There would be two unknowns in just one equation. Percentage abundance for one isotope is needed.
- Radioisotopes occur in nature along with the stable isotopes of an element. The nucleus of a radioisotope is continually altered due to the loss of alpha or beta particles. This in turn affects the atomic mass of that isotope as it undergoes radioactive decay. Changing masses such as this should not be taken into consideration when calculating the average atomic mass of an element. Another reason these isotopes are not used in calculating the average atomic mass of silver is that their natural abundance is usually so small that their contribution to the average atomic mass is insignificant.

# Mhr Chemistry Review Solutions Unit 5

**KJ Lindholm-Leary**



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