

- **Thermodynamics** is the study of energy and its transformations.
 - **First Law of Thermodynamics**: Energy cannot be created or destroyed, only converted from one form to another.

$$\Delta U = Q - W$$

 where ΔU is the change in internal energy, Q is the heat added to the system, and W is the work done by the system.



- **Second Law of Thermodynamics**: The entropy of an isolated system always increases over time.
 - **Entropy** is a measure of the disorder or randomness of a system.
 - **Third Law of Thermodynamics**: The entropy of a perfect crystal at absolute zero is zero.



- **Heat Engines** convert heat into work. The efficiency of a heat engine is given by:

$$\eta = \frac{W}{Q_{in}}$$
 where η is the efficiency, W is the work done, and Q_{in} is the heat input.



- **Refrigerators and Heat Pumps** transfer heat from a cold reservoir to a hot reservoir. The coefficient of performance (COP) is given by:

$$COP = \frac{Q_{out}}{W}$$
 where COP is the coefficient of performance, Q_{out} is the heat removed from the cold reservoir, and W is the work input.



- **Statistical Mechanics** connects the microscopic properties of particles to the macroscopic properties of a system.
 - **Boltzmann's Equation** relates the average energy of a system to its temperature:

$$E = \frac{3}{2} k_B T$$
 where E is the average energy, k_B is Boltzmann's constant, and T is the temperature.



- **Phase Transitions** occur when a substance changes from one state of matter to another (solid, liquid, gas).
 - **Phase Diagrams** show the conditions of temperature and pressure under which different phases of a substance exist.

- **Thermodynamic Potentials** are used to predict the direction of spontaneous processes.
 - **Gibbs Free Energy** is a thermodynamic potential that measures the maximum reversible work that can be performed by a system at constant temperature and pressure.



- **Thermodynamic Tables** provide data for various substances, including their specific heat capacities, enthalpies of formation, and Gibbs free energies of formation.



Microeconomics Final Study Guide

Anthony S. Fauci



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