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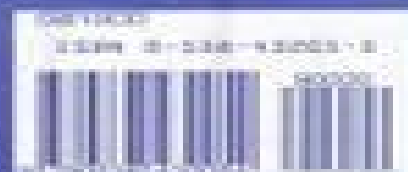
# MEAN JEANS

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A BUSINESS COMMUNITY SIMULATION

## OPERATIONS MANUALS

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# Mean Jeans Manufacturing Co Operations Manuals A Business Community Simulation

**ML Morrison**



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web calculate the mass percent mole fraction molality and molarity of hexane C<sub>6</sub>H<sub>14</sub> D 0.660 g/cm<sup>3</sup> when it is mixed with heptane C<sub>7</sub>H<sub>12</sub> D 0.684 g/cm<sup>3</sup> in a 35.0 : 55.0 ml ratio assume that the volumes are additive

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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0.2885 its mole percent would be 28.85

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web since the mole fraction of the  $\text{O}_2$  is 0.050 it contributes 0.050 or if you will 5.0 of the total pressure thus total pressure is 0.210 0.050 4.2 atm problem 9 a sample of 1.43 g of helium and an unweighed quantity of  $\text{O}_2$  are mixed in a flask at room temperature

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