

Carlos Pintado



Nine Coins/ Nueve monedas

Translated from Spanish by Hilary Vaughn Debel
with an introduction by Richard Blanco

*Winner of The National Poetry Series' Pan Prize for Poetry.
Sponsored by The Center for Writing and Literature at Miami Dade College*

Nine Coins Or Nueve Monedas National Poetry

DJ Losen



Nine Coins Or Nueve Monedas National Poetry:

Whispering the Secrets of Language: An Mental Journey through **Nine Coins Or Nueve Monedas National Poetry**

In a digitally-driven earth wherever screens reign supreme and immediate conversation drowns out the subtleties of language, the profound secrets and mental subtleties concealed within phrases usually get unheard. Yet, situated within the pages of **Nine Coins Or Nueve Monedas National Poetry** a charming literary treasure pulsing with organic emotions, lies an exceptional quest waiting to be undertaken. Written by a skilled wordsmith, this marvelous opus attracts viewers on an introspective trip, lightly unraveling the veiled truths and profound affect resonating within the very cloth of each and every word. Within the psychological depths with this moving evaluation, we shall embark upon a honest exploration of the book is key themes, dissect their fascinating publishing design, and succumb to the powerful resonance it evokes deep within the recesses of readers hearts.

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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2\text{Ag}$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8.3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation
$$8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} = 8.3 \text{ mole } \text{H}_2\text{O}$$

8.3 mole $\text{H}_2\text{S} \times 18 \text{ g } \text{H}_2\text{O} / \text{mole } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$ is your theoretical yield so percent yield
$$\frac{137.1 \text{ g } \text{H}_2\text{O}}{149.4 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$$

problems 6 & 8 are similar to problem 5

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$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$

if the reaction of 21.3 grams of O_2 produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have
$$\frac{21.3 \text{ g } \text{O}_2}{32 \text{ g/mole}} = 0.666 \text{ moles } \text{O}_2$$

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web a write a balanced equation for the reaction
$$2\text{Al} + 3\text{O}_2 \rightarrow \text{Al}_2\text{O}_3$$

b determine the theoretical yield of Al_2O_3
$$\frac{160.0 \text{ g } \text{Al}}{27 \text{ g/mol}} = 5.93 \text{ mol } \text{Al}$$

$$\frac{160.0 \text{ g } \text{Al}}{27 \text{ g/mol}} \times \frac{1 \text{ mol } \text{Al}_2\text{O}_3}{2 \text{ mol } \text{Al}} = 0.296 \text{ mol } \text{Al}_2\text{O}_3$$

$$0.296 \text{ mol } \text{Al}_2\text{O}_3 \times 101.96 \text{ g/mol} = 30.2 \text{ g } \text{Al}_2\text{O}_3$$

c determine the percent yield
$$\frac{26.98 \text{ g } \text{Al}_2\text{O}_3}{30.2 \text{ g } \text{Al}_2\text{O}_3} \times 100 = 89.3\%$$

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