

Name: _____

Class: _____

OXIDATION NUMBERS WORKSHEET

The *oxidation number* of an atom is the apparent charge assigned to it in a particular molecule, ion or compound. Certain rules are followed in assigning oxidation numbers.

Use rules at the right to assign oxidation numbers to each element in each of the given formulas.

Example

	OXIDATION NO.	RULES USED
H_2O	<u>H=+1, O= -2</u>	<u>6 & 7</u>
N_2	<u>N= 0</u>	<u>1</u>

Rules The oxidation number of:

1. A free element in the uncombined state is 0.
2. A monatomic ion equals the charge on the ion.
3. Oxygen in a compound is generally -2
4. Hydrogen in a compound is +1
5. In a neutral compound the sum of the oxidation numbers for all the atoms is 0.
6. In a polyatomic ion the sum of the oxidation numbers for all the atoms is equals the charge on the ion.
7. Group 1 metals is +1.
8. Group 2 metals is +2.

		RULES USED
1.	Cl_2	_____
2.	Cl^-	_____
3.	Na	_____
4.	Na^+	_____
5.	KCl	_____
6.	H_2S	_____
7.	CaO	_____
8.	H_2SO_4	_____
9.	NO_3^-	_____
10.	$\text{Cr}_2\text{O}_7^{2-}$	_____
11.	NH_4Cl	_____
12.	NH_3	_____
13.	NO_2	_____
14.	CaH_2 (calcium hydride)	_____
15.	Na_2O_2 (sodium peroxide)	_____

Oxidation Number Practice Answer Key

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Table of Contents Oxidation Number Practice Answer Key

1. Understanding the eBook Oxidation Number Practice Answer Key
 - The Rise of Digital Reading Oxidation Number Practice Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Oxidation Number Practice Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Oxidation Number Practice Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Oxidation Number Practice Answer Key
 - Personalized Recommendations
 - Oxidation Number Practice Answer Key User Reviews and Ratings
 - Oxidation Number Practice Answer Key and Bestseller Lists

5. Accessing Oxidation Number Practice Answer Key Free and Paid eBooks
 - Oxidation Number Practice Answer Key Public Domain eBooks
 - Oxidation Number Practice Answer Key eBook Subscription Services
 - Oxidation Number Practice Answer Key Budget-Friendly Options
6. Navigating Oxidation Number Practice Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Oxidation Number Practice Answer Key Compatibility with Devices
 - Oxidation Number Practice Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Oxidation Number Practice Answer Key
 - Highlighting and Note-Taking Oxidation Number Practice Answer Key
 - Interactive Elements Oxidation Number Practice Answer Key
8. Staying Engaged with Oxidation Number Practice Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Oxidation Number Practice Answer Key
9. Balancing eBooks and Physical Books Oxidation Number Practice Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Oxidation Number Practice Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Oxidation Number Practice Answer Key
 - Setting Reading Goals Oxidation Number Practice Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Oxidation Number Practice Answer Key
 - Fact-Checking eBook Content of Oxidation Number Practice Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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