

Open Water Diver Course Final Exam Answers/ PADI Open Water Diver Final Exam Answers Updated 2023-2023 Latest Version

1) If an object is neutrally buoyant (does not sink or float) in fresh water, the same object placed into salt water would

- A sink.
- B either sink or float.
- C do nothing.
- D float. ----- Correct Answer ----- D float.

Salt water is heavier than fresh water because it has dissolved minerals in it. This means it causes more upward force (buoyancy) on an object. An object that is neutrally buoyant in fresh water would float in salt water because there is greater upward force.

See Being a Diver I - Buoyancy and Controlling Buoyancy.

2) I blow up a balloon, tie it off, and take it to the bottom of the swimming pool. What will happen to the balloon and the air inside it?

- A The balloon will get bigger and the air inside the balloon will be less dense (molecules move further apart).
- B The balloon will get smaller and the air inside the balloon will be more dense (molecules move closer together).
- C The balloon will get bigger and the air inside the balloon will be more dense.
- D The balloon will get smaller and the air inside the balloon will be less dense. ----- Correct Answer ----- B The balloon will get smaller and the air inside the balloon will be more dense (molecules move closer together).

As water pressure increases, the volume of an air space will decrease. This causes the density of the air inside to increase and air molecules are pushed closer together. The balloon would get smaller and the air density inside would be greater.

See Being a Diver I - Water Pressure and Air Volume Effects.

3) I turn a glass upside down, trap the air in it by putting it in water, and then I take the glass down to 10 meters. The air space would

- A become half the size it was at the surface.
- B not change in size.
- C become 1/3 the size it was at the surface.

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web he concluded that at a constant pressure the volume of the gas would expand at a particular rate for each degree of temperature rise that being 1 267 per degree in 1874 victor regnault obtained better experimental results showing this number to be 1 273 and concluded that the pressure would approach zero at 1 273 degrees c

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web the volume of 1.00 mol of any gas at stp standard temperature 273.15 K and pressure 1 atm is measured to be 22.414 L we can substitute 101.325 kPa for pressure 22.414 L for volume and 273.15 K for temperature into the ideal gas equation and solve for r

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web chemists sometimes make comparisons against a standard temperature and pressure stp for reporting properties of gases 273.15 K and 1 atm 101.325 kPa 1 atm stp one mole of an ideal gas has a volume of about 22.4 L this is referred to as the standard molar volume figure 9.18

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