

**Table K
Selected Polyatomic Ions**

Formula	Name	Formula	Name
ClO_4^-	perchlorate	ClO_2^-	chlorite
ClO_3^-	chlorate	ClO_2^+	dichlorine oxide
NH_4^+	ammonium	MnO_4^-	permanganate
$\text{C}_2\text{H}_3\text{O}_2^-$ CH_3COO^-	acetate	NO_2^-	nitrite
CN^-	cyanide	NO_3^-	nitrate
CO_3^{2-}	carbonate	O_2^{2-}	peroxide
HCO_3^-	hydrogen carbonate	OH^-	hydroxide
$\text{C}_2\text{O}_4^{2-}$	oxalate	PO_4^{3-}	phosphate
CrO_4^{2-}	chromate	SCN^-	thiocyanate
CrO_2^-	chromite	SO_3^{2-}	sulfite
CrO_4^{2-}	chromate	SO_3^{2-}	sulfite
$\text{Cr}_2\text{O}_7^{2-}$	dichromate	HSO_4^-	hydrogen sulfate
Fe_3O_4	pyrrole	$\text{S}_2\text{O}_8^{2-}$	persulfate

**Table L
Solubility Guidelines for Aqueous Solutions**

Basic Most Forms Soluble Compounds	Exceptions	Basic Most Forms Soluble Compounds*	Exceptions
Group 1 ions (Li^+ , Na^+ , etc.)		acetate (CH_3CO_2^-)	when combined with Group 1 ions or ammonium (NH_4^+)
ammonium (NH_4^+)		chromate (CrO_4^{2-})	when combined with Group 1 ions, Ca^{2+} , Sr^{2+} , or ammonium (NH_4^+)
nitrate (NO_3^-)		phosphate (PO_4^{3-})	when combined with Group 1 ions or ammonium (NH_4^+)
hydrogen carbonate (HCO_3^-)		oxide (O^{2-})	when combined with Group 1 ions or ammonium (NH_4^+)
halides (F^- , Cl^-)		hydroxide (OH^-)	when combined with Group 1 ions, Ca^{2+} , Sr^{2+} , Ba^{2+} , or ammonium (NH_4^+)
halides (Br^- , I^-)	when combined with Ag^+ , Pb^{2+} , or Hg_2^{2+}		
sulfates (SO_4^{2-})	when combined with Ag^+ , Ca^{2+} , Sr^{2+} , Ba^{2+} , or Pb^{2+}		

*Transition metal ions may have solubility as H_2O

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Table of Contents New York State Ratings Guide Chemistry Regents

1. Understanding the eBook New York State Ratings Guide Chemistry Regents
 - The Rise of Digital Reading New York State Ratings Guide Chemistry Regents
 - Advantages of eBooks Over Traditional Books
2. Identifying New York State Ratings Guide Chemistry Regents
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an New York State Ratings Guide Chemistry Regents
 - User-Friendly Interface
4. Exploring eBook Recommendations from New York State Ratings Guide Chemistry Regents
 - Personalized Recommendations
 - New York State Ratings Guide Chemistry Regents User Reviews and Ratings

- New York State Ratings Guide Chemistry Regents and Bestseller Lists
- 5. Accessing New York State Ratings Guide Chemistry Regents Free and Paid eBooks
 - New York State Ratings Guide Chemistry Regents Public Domain eBooks
 - New York State Ratings Guide Chemistry Regents eBook Subscription Services
 - New York State Ratings Guide Chemistry Regents Budget-Friendly Options
- 6. Navigating New York State Ratings Guide Chemistry Regents eBook Formats
 - ePub, PDF, MOBI, and More
 - New York State Ratings Guide Chemistry Regents Compatibility with Devices
 - New York State Ratings Guide Chemistry Regents Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of New York State Ratings Guide Chemistry Regents
 - Highlighting and Note-Taking New York State Ratings Guide Chemistry Regents
 - Interactive Elements New York State Ratings Guide Chemistry Regents
- 8. Staying Engaged with New York State Ratings Guide Chemistry Regents
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers New York State Ratings Guide Chemistry Regents
- 9. Balancing eBooks and Physical Books New York State Ratings Guide Chemistry Regents
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection New York State Ratings Guide Chemistry Regents
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine New York State Ratings Guide Chemistry Regents
 - Setting Reading Goals New York State Ratings Guide Chemistry Regents
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of New York State Ratings Guide Chemistry Regents
 - Fact-Checking eBook Content of New York State Ratings Guide Chemistry Regents
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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