

Organic Reactions Summary Sheet

Reactants	Products	Conditions	Classification	Example	Uses /Other
Alkanes					
+ Oxygen	Complete: CO ₂ and water Incomplete: CO and water	Little bit of heat to get it going and plenty of oxygen Little bit of heat to get it going, and not enough O ₂ , oxygen is limiting reactant	Combustion (also redox reactions)	$2 \text{C}_8\text{H}_{18} + 13 \text{O}_2 \rightarrow 8 \text{CO}_2 + 10 \text{H}_2\text{O}$ $2 \text{C}_8\text{H}_{18} + 9 \text{O}_2 \rightarrow 8 \text{CO} + 10 \text{H}_2\text{O}$	CO harmful to humans, unburned hydrocarbons produce smog
+ Halogen (F ₂ , Cl ₂ , Br ₂ , I ₂)	haloalkane	UV light, hv	Substitution		
Alkenes					
+ halogen (F ₂ , Cl ₂ , Br ₂ , I ₂)	Dihaloalkane		Addition		Uses: makes starting material for PVC and indication of alkenes Br ₂ dissolved in dichloromethane is used to distinguish between alkenes and alkanes. If reddish-brown color of Br ₂ disappears when added to unknown, the unknown has alkenes in it.
+ hydrogen	Alkane	Platinum or palladium catalyst	Addition (also called reduction)		Hydrogenation of oil for margarine
+ acid halide (HF, HCl, HBr, HI)	Haloalkane		Addition		Haloalkanes are useful starting materials for organic syntheses
+ water	Alcohol	Acid conditions	Addition		commercial manufacture of ethanol
Alcohols					
+ Carboxylic acid	Ester + water	Acidic solution and heat	Condensation		Esters are used in flavouring agents, plasticizers, as solvents in perfume, polyester
+ oxidizing agent (KMnO ₄ or K ₂ Cr ₂ O ₇)	Primary: Aldehyde	For incomplete oxidation: Distillation in acidic aqueous soln	Oxidation		

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