

Reactions of Amines and Epoxides

Preparation of Amines

Amination of alcohols: Polyalkylation is a serious problem for indirect synthesis of amines



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Bond Types, Strength, Length, Polarity and Dipole Moment

Two different bond types

There are two types of bonding: **ionic bonding** and **covalent bonding**

- **Ionic bonds** are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron).



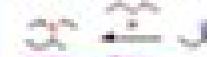
- **Covalent bonds** are formed by sharing of electrons between two atoms.



The Relative Electronegativity of Atoms



Bonded and Unbonded electron pairs



Bonding electrons



There are two main categories

Definition of electron pair



Sigma (σ) and pi (π) bonds

Covalent bonds are classified in **sigma (σ) bonds** and **pi (π) bonds**

- **Sigma bonds** are formed by head-on (frontal) overlap of orbitals.
- **pi bonds** are formed by side-on overlap of orbitals.
- **Any single bond** is a **sigma bond**.
- **Any multiple bond** consists of one **sigma** and **n pi** bonds.



σ bonds are stronger than π bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds can be **polar** and **nonpolar**.



- If the difference in electronegativity is **less than 0.5**, the electrons are shared equally (nonpolar bond).
- If the difference in electronegativity is **between 0.5 and 1.7**, the electrons are shared unequally (polar bond).
- If the difference in electronegativity is **greater than 1.7**, the electrons are transferred (ionic bond).

- A difference of **1.7 or higher** can be high, but the difference can be higher (shared) and a lower bond is formed.

Dipole Moment

The polarity of a bond is described by the **dipole moment (μ)**.



- The dipole moment is defined as a vector pointing from the less electronegative to the more electronegative atom.
- The more electronegative atom has a partial negative charge (δ⁻) and the less electronegative atom has a partial positive charge (δ⁺).
- **δ⁻** (delta minus) indicates an electron deficient atom.
- **δ⁺** (delta plus) indicates an electron rich atom.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipoles is used.

The geometry of the molecule is important. Symmetrical molecules are nonpolar even if the dipoles are present. A good example is carbon dioxide.



Consider the vector sum of the dipoles pointing from the center of the molecule.

Bond Length and Bond Strength

The longer the bond, the weaker the bond.

Bond	H-H	H-Cl	H-Br	H-I
Length	109 pm	127 pm	141 pm	154 pm
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

Double bonds are stronger than single bonds, and triple bonds are stronger than double bonds.

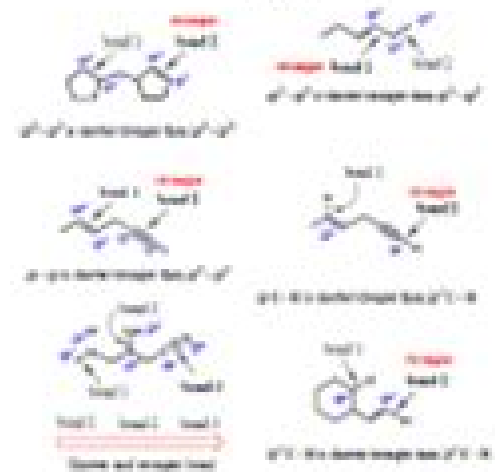


The longer the bond, the weaker the bond.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp ³ -C-C	sp ² -C-C	sp-C-C
Hybridization	sp ³ -sp ³	sp ² -sp ²	sp-sp
Length	154 pm	134 pm	121 pm
Strength	376 kJ/mol	611 kJ/mol	591 kJ/mol
Hybridization	sp ³ -sp ²	sp ² -sp	sp-sp
Length	147 pm	133 pm	120 pm
Strength	410 kJ/mol	602 kJ/mol	591 kJ/mol

Higher s character = stronger and shorter bond

Examples



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Peterson's



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