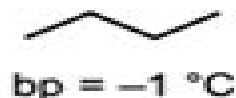
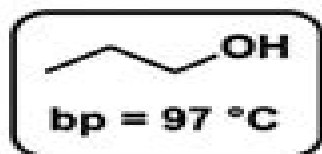
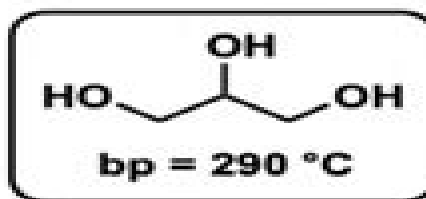


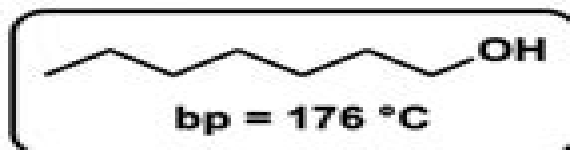
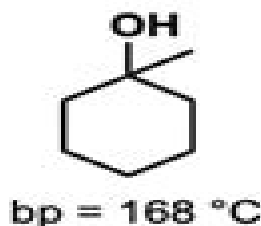
Three pairs of compounds with similar molecular weight are shown below. Choose the compound with the **higher boiling point and explain**.



**Hydrogen bonding**, present in alcohols but not hydrocarbons, leads to strong intermolecular forces and increases the boiling point significantly.



Glycerol has **3 OH groups**, which lead to a much more extensive hydrogen bonding network and higher boiling point compared to the 1 OH group in *n*-pentanol.



1-methylcyclohexanol is a tertiary alcohol and has more steric crowding around the alcohol. *n*-Heptanol has **less steric hindrance** around the OH group, making it more available for hydrogen bonding and leading to a slightly higher boiling point.

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