

1. What is Resonance?

The true distribution of electrons in a molecule cannot always be accurately depicted through chemical formulae. Resonance is a tool to show this.

So what about applying this concept to these molecules?



In reality, the charge in both of these molecules is distributed equally between the different units.

Important: this is not an equilibrium!



Instead, the "true" structure of the molecule is a hybrid of these two forms.



2. How to Use Curved Arrows

There are only three legal "moves" we can do with curved arrows to show resonance.



3. When Resonance Forms are Not Equal: How to tell which are most stable?

Rule #1 - The Fewer Charges, the Better

The more stable a resonance form is, the more it will contribute to the resonance hybrid.



Rule #2 - Full Orbits Are Preferred

A resonance form in which all atoms have a full orbit is more significant than one where at least one atom lacks a full orbit.



Rule #3 - Place a negative charge on the most electronegative atom



If given a choice, the best place to put a negative charge on a resonance form is on the **least electronegative** atom.

Use acidity trends to determine the "best" resonance form in these cases:

- Electronegativity
- Polarizability
- Induction
- Orbital hybridization

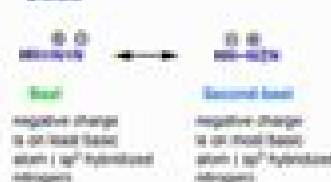
Polarizability



Electron withdrawing groups



Orbitals



Rule #4 Stabilizing positive charge

If the resonance form has a positive charge, try and find a resonance form in which all orbits are full (Rule #2).

If all orbits are not full, place positive charge on the atom best able to support a carbocation i.e. the most substituted carbon.



Rule #5 - Aromaticity (Org 2 Topic)

A resonance form that is aromatic will have a disproportionate contribution to the resonance hybrid.



4. Applications of Resonance

π \rightarrow donation

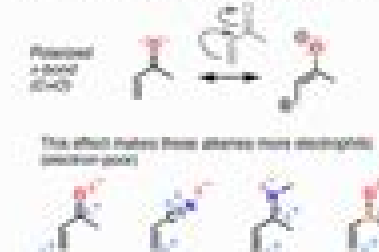
Atoms attached to an atom with a lone pair such as O, N, S, etc. often called "heteroatoms" have an important resonance form with a negative charge adjacent to the carbon atom bond.

Example



π \rightarrow accepting

When double bonds are connected to a polarized π bond, the molecule will have a significant resonance form where there is positive charge on the adjacent carbon.



Omissions, Mistakes, Suggestions?

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Organic Chemistry Semester 1 Mechanism Study Guide

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