

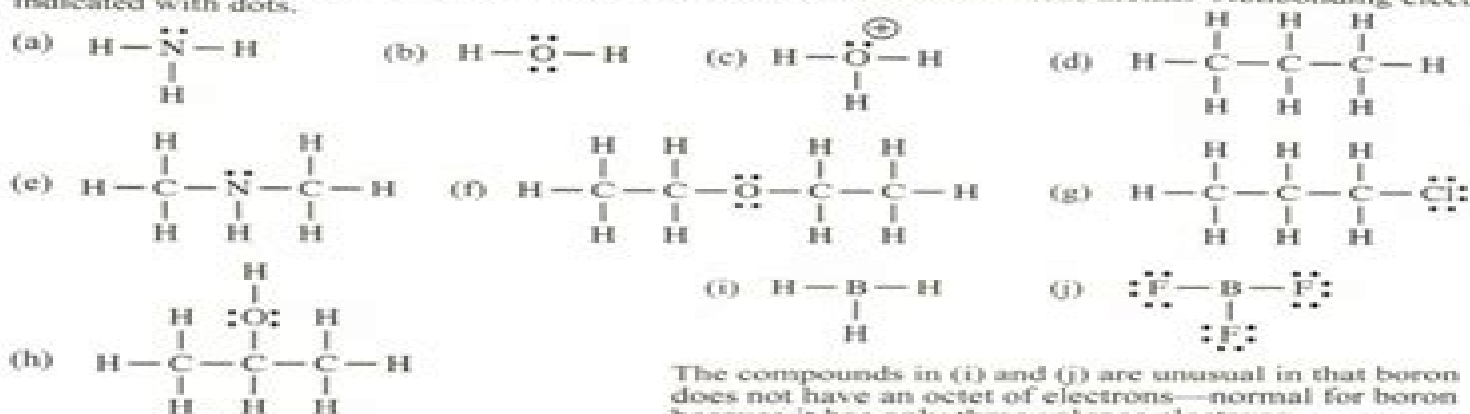
CHAPTER 1—STRUCTURE AND BONDING

1-1

(a) Nitrogen has atomic number 7, so all nitrogen atoms have 7 protons. The mass number is the total number of neutrons and protons; therefore, ^{13}N has 6 neutrons, ^{14}N has 7 neutrons, ^{15}N has 8 neutrons, ^{16}N has 9 neutrons, and ^{17}N has 10 neutrons.

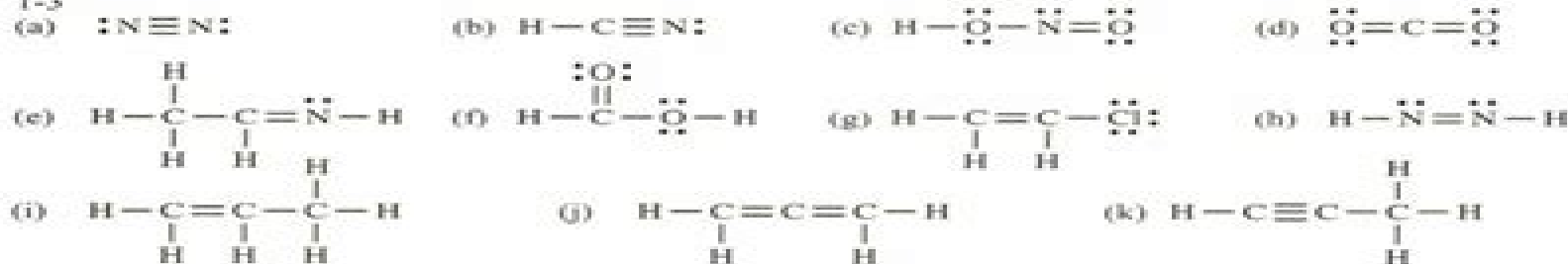
(b)	Na	$1s^2 2s^2 2p^6 3s^1$	P	$1s^2 2s^2 2p^6 3s^2 3p_x^1 3p_y^1 3p_z^1$
	Mg	$1s^2 2s^2 2p^6 3s^2$	S	$1s^2 2s^2 2p^6 3s^2 3p_x^2 3p_y^1 3p_z^1$
	Al	$1s^2 2s^2 2p^6 3s^2 3p_x^1$	Cl	$1s^2 2s^2 2p^6 3s^2 3p_x^2 3p_y^2 3p_z^1$
	Si	$1s^2 2s^2 2p^6 3s^2 3p_x^1 3p_y^1$	Ar	$1s^2 2s^2 2p^6 3s^2 3p_x^2 3p_y^2 3p_z^2$

1-2 Lines between atom symbols represent covalent bonds between those atoms. Nonbonding electrons are indicated with dots.



The compounds in (i) and (j) are unusual in that boron does not have an octet of electrons—normal for boron because it has only three valence electrons.

1-3



1-4 There are no unshared electron pairs in parts (i), (j), and (k).



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