

POLARITY REVIEW

1. Molarity is a type of _____, similar to g/L , mg/mL , cm^3/cm^3 , etc. etc. etc.
2. Molarity is defined as M of _____ dissolved within one _____ of solution.
3. M is the unit for molarity, but the unit for molarity can also be written as:
4. Analyze the images below and identify the molarity of the solutions pictured.

2 moles NaCl dissolved in 2 liters of solution



_____ M

4.0 moles NaCl dissolved in 1.000 L of solution



_____ M

4 moles NaCl dissolved in 0.5 L of solution



_____ M

0.5 moles NaCl dissolved in 0.5 liters of solution



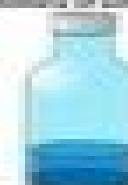
_____ M

2 moles NaCl dissolved in 0.25 L of solution



_____ M

1 mole NaCl dissolved in 0.500 L of solution



_____ M

5. Analyze the images below and identify how many total moles of the compound are present

1 L of 0.500 solution



_____ moles

1000 mL of 1.000 solution



_____ moles

500 mL of 1.000 solution



_____ moles

0.5 L of 1.000 solution



_____ moles

2000 mL of 0.5 solution



_____ moles

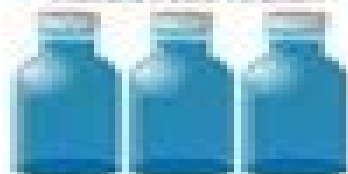
0.25 L of 1.000 solution



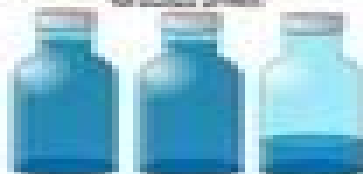
_____ moles

6. Analyze the images below and answer the prompts accordingly. Be sure to include the correct units and identity of the solution on question. Ex: 2 moles CH_4 or 3M MgO

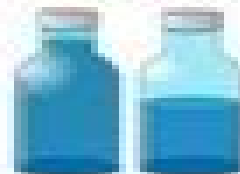
How many moles of NaCl are in 1000 mL of a 0.500 solution?



What is the concentration in molarity of 2000 mL of a solution containing 10 moles of NaCl ?



How many moles of NaCl are in 0.5 L of a 0.500 solution?



Answer the below 3a and 3b with the number of grams. 3a. How many grams of sugar are in 1000 mL of Diet Coke? 3b. How many grams of sugar are in 100 mL of a 10% solution of sugar in 1000 mL of water? (What is the density of water? NaCl is 2.16 g/mL)



Molarity Concept Review

**Barron's Educational Series, Neil D.
Jespersen, Pamela Kerrigan**



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