

Masses and Moles

Look at this example of how to calculate the mass from the number of moles:

Mass of substance = number of moles \times mass of 1 mole

Which is the same as saying....

Mass = number of moles \times Relative Formula Mass

$$m = n \times M_r$$



Example

What is the mass of 2 moles of carbon?

m	$=$	$n \times M_r$
	$=$	2×12
	$=$	$24g$

Use the Periodic Table in the front of the book to answer these questions:

Q1 What is the mass of...?

- | | |
|--|--|
| <ul style="list-style-type: none"> a) 1 mole of carbon atoms b) 2 moles of sodium atoms c) 2 moles of aluminium atoms d) 3 moles of lithium atoms e) 10 moles of iodine atoms | <ul style="list-style-type: none"> f) 1 mole of oxygen molecules g) 2 moles of chlorine molecules h) 1 mole of nitrogen molecules i) 0.5 moles of nitrogen molecules j) 0.1 moles of carbon molecules |
|--|--|

Q2 What is the mass of...?

- | | |
|--|--|
| <ul style="list-style-type: none"> a) 1 mole of carbon dioxide (CO_2) b) 5 moles of carbon dioxide (CO_2) c) 3 moles of water (H_2O) d) 2 moles of sulphur trioxide (SO_3) e) 3 moles of nitric acid (HNO_3) | <ul style="list-style-type: none"> f) 10 moles of hydrochloric acid (HCl) g) 100 moles of sulphuric acid (H_2SO_4) h) 50 moles of sodium hydroxide (NaOH) i) 30 moles of calcium oxide (CaO) j) 25 moles of sodium chloride (NaCl) |
|--|--|

Q3 What is the mass of...?

- | | |
|--|--|
| <ul style="list-style-type: none"> a) 0.1 moles of copper oxide (CuO) b) 0.1 moles of calcium carbonate (CaCO_3) c) 0.01 moles of sulphuric acid (H_2SO_4) d) 0.2 moles of calcium chloride (CaCl_2) e) 0.25 moles of sodium carbonate (Na_2CO_3) | <ul style="list-style-type: none"> f) 0.05 moles of sodium hydroxide (NaOH) g) 0.58 moles of aluminium oxide (Al_2O_3) h) 0.1 moles of propanoic acid ($\text{C}_2\text{H}_5\text{COOH}$) i) 0.5 moles of potassium manganate(VII) (KMnO_4) j) 0.25 moles of Hydrated Magnesium Sulphate ($\text{MgSO}_4 \cdot 7\text{H}_2\text{O}$) |
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