



Figure 1.1: NERC Certification Process Flowchart

Nerc Certification Study Guide

Raffaela Di Napoli

A red circular graphic with a gradient, appearing as a partial circle or a thick arc, located to the right of the author's name.

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 $\text{H}_2\text{O}(\text{g}) \rightarrow \frac{1}{2} \text{O}_2(\text{g}) + \text{H}_2(\text{g}) \quad \Delta H = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) \rightarrow 2 \text{O}(\text{g}) + 3 \text{H}_2(\text{g}) \quad \Delta H = 283 \text{ kJ}$

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 $\text{C}_2\text{H}_6(\text{g}) + 7 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{l}) \quad \Delta H = 1560 \text{ kJ/mole}$
 $\text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l}) \quad \Delta H = 285 \text{ kJ/mole}$

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web thus for reaction 9 2 in which 1 mole of H_3O^+ and 1 mole of OH^- combine to form 2 moles of H_2O ΔH 55.8 kJ as shown note that the enthalpy change $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta H_{\text{products}}$

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