

Module 1

Financial Accounting for SMEs

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|----------|--|
| QUESTION | <p>1. Explain the difference between a <i>strong</i> and a <i>weak</i> acid. Give examples of each.</p> <p>2. Write the chemical formula for the conjugate base of each of the following acids:</p> <p>3. Calculate the pH of a 0.1 M solution of acetic acid ($K_a = 1.8 \times 10^{-5}$).</p> |
| ANSWER | <p>1. A strong acid is one that completely dissociates in water, while a weak acid only partially dissociates. Examples of strong acids are HCl, HNO₃, and H₂SO₄. Examples of weak acids are acetic acid (CH₃COOH) and carbonic acid (H₂CO₃).</p> <p>2. The conjugate base of an acid is formed by removing a proton (H⁺) from the acid. For example, the conjugate base of HCl is Cl⁻, and the conjugate base of CH₃COOH is CH₃COO⁻.</p> <p>3. The pH of a 0.1 M solution of acetic acid can be calculated using the K_a value. The equation for the dissociation of acetic acid is:</p> $\text{CH}_3\text{COOH} \rightleftharpoons \text{CH}_3\text{COO}^- + \text{H}^+$ <p>The K_a expression is:</p> $K_a = \frac{[\text{CH}_3\text{COO}^-][\text{H}^+]}{[\text{CH}_3\text{COOH}]}$ <p>Let x be the concentration of H^+ ions at equilibrium. Then:</p> $1.8 \times 10^{-5} = \frac{x^2}{0.1 - x}$ <p>Since x is small compared to 0.1, we can approximate $0.1 - x \approx 0.1$:</p> $1.8 \times 10^{-5} \approx \frac{x^2}{0.1}$ $x^2 \approx 1.8 \times 10^{-6}$ $x \approx 1.34 \times 10^{-3}$ <p>Therefore, the pH of the solution is:</p> $\text{pH} = -\log[\text{H}^+] = -\log(1.34 \times 10^{-3}) \approx 2.87$ |
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**Peter D. Easton, John J. Wild, Robert F.
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