

Molarity = Moles of solute / Liters of Solution (abbreviation = M)

Molality = Moles of solute / Kg of Solvent (abbreviation = m)

Normality = number of equivalent of solute x Molarity of Solution (abbreviation = N)

Mass Percent = mass of solute / mass of solution

1. How many moles of ethyl alcohol, C_2H_5OH , are present in 65 mL of a 1.5 M solution?

$$\frac{65 \text{ mL } C_2H_5OH}{1000 \text{ mL } C_2H_5OH} \times \frac{1 \text{ L } C_2H_5OH}{1 \text{ L } C_2H_5OH} \times \frac{1.5 \text{ mole } C_2H_5OH}{1 \text{ L } C_2H_5OH} = 0.098 \text{ moles } C_2H_5OH$$

2. How many liters of a 6.0 M solution of acetic acid CH_3COOH , contain 0.0030 moles of acetic acid?

$$\frac{0.0030 \text{ moles } CH_3COOH}{6.0 \text{ moles } CH_3COOH} \times \frac{1 \text{ L solution}}{1 \text{ L solution}} = 5.0 \times 10^{-4} \text{ L solution}$$

3. You want 85 g of KOH. How much of a 3.0 m solution of KOH will provide it?

$$\frac{85 \text{ g KOH}}{56.11 \text{ g KOH}} \times \frac{1 \text{ mole}}{1 \text{ mole}} \times \frac{1 \text{ kg solvent}}{3.0 \text{ mole KOH}} \times \frac{1000 \text{ g solvent}}{1 \text{ kg solvent}} = 5.0 \times 10^2 \text{ g solvent}$$

$$5.0 \times 10^2 \text{ g solvent} + 85 \text{ g KOH} = 590 \text{ g solution}$$

4. If you dissolve 0.70 moles of HCl in enough water to prepare 250 mL of solution, what is the molarity of the solution you have prepared?

$$M = \frac{0.70 \text{ moles HCl}}{0.250 \text{ L solution}} = 2.8 \text{ M HCl}$$

5. A solution is prepared by adding 2.0 L of 6.0 M HCl to 500 mL of a 9.0 M HCl solution. What is the molarity of the new solution? (Remember, the volumes are additive)

$$\frac{2.0 \text{ L solution}}{1 \text{ L solution}} \times \frac{6.0 \text{ moles HCl}}{1 \text{ L solution}} = 12 \text{ moles HCl}$$

$$\frac{0.500 \text{ L solution}}{1 \text{ L solution}} \times \frac{9.0 \text{ moles HCl}}{1 \text{ L solution}} = 4.5 \text{ moles HCl}$$

$$12 \text{ moles HCl} + 4.5 \text{ moles HCl} = 17 \text{ moles}$$

$$2.0 \text{ L solution} + 0.500 \text{ L solution} = 2.5 \text{ L solution}$$

$$M = \frac{17 \text{ moles}}{2.5 \text{ L solution}} = 6.8 \text{ M HCl}$$

6. Convert the following Molarities to Normalities.

a. 2.5 M HCl = 2.5 N

b. 1.4 M H_2SO_4 = 2.8 N

c. 1.0 M NaOH = 1.0 N

d. 0.5 M $Ca(OH)_2$ = 1 N

Molarity Practice Answer Key Chemfiesta

Tobias Bleicker

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Molarity Practice Answer Key Chemfiesta:

The book delves into Molarity Practice Answer Key Chemfiesta. Molarity Practice Answer Key Chemfiesta is an essential topic that needs to be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Molarity Practice Answer Key Chemfiesta, encompassing both the fundamentals and more intricate discussions.

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- Chapter 1: Introduction to Molarity Practice Answer Key Chemfiesta
- Chapter 2: Essential Elements of Molarity Practice Answer Key Chemfiesta
- Chapter 3: Molarity Practice Answer Key Chemfiesta in Everyday Life
- Chapter 4: Molarity Practice Answer Key Chemfiesta in Specific Contexts
- Chapter 5: Conclusion

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