

# The Mole Study Guide

**Mass Composition:** What's your element's % of mass after the reaction? Molar mass to obtain your value!

10. There are 3.75 moles of  $\text{CO}_2$ .

11. There are 3.75 moles of  $\text{CO}_2$ .

12. There are 3.75 moles of  $\text{CO}_2$ .

13. There are 3.75 moles of  $\text{CO}_2$ .

14. There are 3.75 moles of  $\text{CO}_2$ .

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21. There are 3.75 moles of  $\text{CO}_2$ .

22. There are 3.75 moles of  $\text{CO}_2$ .

23. There are 3.75 moles of  $\text{CO}_2$ .

24. There are 3.75 moles of  $\text{CO}_2$ .

25. There are 3.75 moles of  $\text{CO}_2$ .

26. There are 3.75 moles of  $\text{CO}_2$ .

Name \_\_\_\_\_ Period \_\_\_\_\_ Date \_\_\_\_\_

## The Mole Study Guide

Fill in the Blank boxes for the image below.



Write the Empirical Formula for each compound. Use subscripts to show the relative ratio of atoms.

Molecular	Empirical
$\text{C}_6\text{H}_{12}\text{O}_6$	
$\text{C}_4\text{H}_8\text{O}_4$	
$\text{C}_8\text{H}_{16}\text{O}_8$	
$\text{C}_{12}\text{H}_{24}\text{O}_{12}$	
$\text{C}_{16}\text{H}_{32}\text{O}_{16}$	

Calculate the Molar Mass for the following compounds.

10.  $\text{H}_2\text{O}$

11.  $\text{H}_2\text{O}_2$

12.  $\text{H}_2\text{O}$

13.  $\text{H}_2\text{O}_2$

Calculate the Percent Composition for each element in the following compound.

10.  $\text{H}_2\text{O}$

11.  $\text{H}_2\text{O}_2$

12.  $\text{H}_2\text{O}$

13.  $\text{H}_2\text{O}_2$

14.  $\text{H}_2\text{O}$

15.  $\text{H}_2\text{O}_2$

16.  $\text{H}_2\text{O}$

17.  $\text{H}_2\text{O}_2$

18.  $\text{H}_2\text{O}$

19.  $\text{H}_2\text{O}_2$

Determine the Empirical and Molecular Formula for the compound given the information below.

10. A compound has the following percent composition: 58.80% Hg, 21.80% Cl, 1.40% N. The molar mass for Aqueous is 100.0 g/mol.

a. If you have 100 g of Aqueous, how many grams of Hg, Cl, and N do you have?

b. Determine the Empirical Formula for Aqueous.

c. Determine the Molecular Formula for Aqueous.

# Mole Study Guide

**James E. Brady**



## **Mole Study Guide:**

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