

## Chemistry Chapter 14a - Thermodynamics

- Be able to predict the spontaneity of a reaction from either its standard Gibbs free energy or its standard enthalpy and entropy values (what criteria to look at for spontaneity)
- Remember that following relationship always goes with your energy:

$$\Delta G = \Delta H - T\Delta S \text{ (Always use this equation for Gibbs)}$$

Use the following approximate thermodynamic quantities given in appendices to predict the spontaneity of a reaction using the equation  $\Delta G = \Delta H - T\Delta S$  (remember that  $\Delta G$  must be negative for a spontaneous reaction).  
 Values generally accepted given based on finding a representative thermodynamic value.

Water	$\Delta_f H^\circ = -285.8 \text{ kJ/mol}$
Hydrogen gas	$\Delta_f H^\circ = 0 \text{ kJ/mol}$
Oxygen gas	$\Delta_f H^\circ = 0 \text{ kJ/mol}$
Carbon	$\Delta_f H^\circ = 0 \text{ kJ/mol}$
Hydrogen peroxide (l)	$\Delta_f H^\circ = -187.8 \text{ kJ/mol}$
Hydrogen peroxide (g)	$\Delta_f H^\circ = -136.1 \text{ kJ/mol}$
Carbon monoxide	$\Delta_f H^\circ = -110.5 \text{ kJ/mol}$
Carbon dioxide	$\Delta_f H^\circ = -393.5 \text{ kJ/mol}$
Standard entropy	$\Delta_f S^\circ = 167.2 \text{ J/mol}\cdot\text{K}$

Hydrogen gas	$\Delta_f S^\circ = 130.7 \text{ J/mol}\cdot\text{K}$
Oxygen gas	$\Delta_f S^\circ = 205.1 \text{ J/mol}\cdot\text{K}$
Water	$\Delta_f S^\circ = 69.9 \text{ J/mol}\cdot\text{K}$
Carbon	$\Delta_f S^\circ = 5.7 \text{ J/mol}\cdot\text{K}$
Hydrogen peroxide (l)	$\Delta_f S^\circ = 109.1 \text{ J/mol}\cdot\text{K}$
Hydrogen peroxide (g)	$\Delta_f S^\circ = 151.9 \text{ J/mol}\cdot\text{K}$
Carbon monoxide	$\Delta_f S^\circ = 191.6 \text{ J/mol}\cdot\text{K}$
Carbon dioxide	$\Delta_f S^\circ = 213.8 \text{ J/mol}\cdot\text{K}$

### Step 1: $\Delta G$

### Step 2: Concentration

Use the following concentrations to get  $\Delta G$  from  $\Delta G^\circ$  for a reaction involving gas or liquid reactants:

- Hydrogen gas concentration = 1 atm
- Oxygen gas concentration = 1 atm
- Water concentration = 1 M
- Carbon concentration = 1 M
- Hydrogen peroxide (l) concentration = 1 M
- Hydrogen peroxide (g) concentration = 1 atm
- Carbon monoxide concentration = 1 atm
- Carbon dioxide concentration = 1 atm

- Hydrogen gas concentration = 1 atm
- Oxygen gas concentration = 1 atm

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**Michal Rosen-Zvi**



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