

# CHAPTER 10 The Mole

## Section 1 Measuring Matter

pages 320–324

### Practice Problems

pages 323–324

1. Zinc (Zn) is used to form a corrosion-inhibiting surface on galvanized steel. Determine the number of Zn atoms in 2.50 mol of Zn.

$$2.50 \text{ mol Zn} \times \frac{6.02 \times 10^{23} \text{ atoms Zn}}{1 \text{ mol Zn}} \\ = 1.51 \times 10^{24} \text{ atoms of Zn}$$

2. Calculate the number of molecules in 11.5 mol of water (H<sub>2</sub>O).

$$11.5 \text{ mol H}_2\text{O} \times \frac{6.02 \times 10^{23} \text{ molecules H}_2\text{O}}{1 \text{ mol H}_2\text{O}} \\ = 6.92 \times 10^{24} \text{ molecules H}_2\text{O}$$

3. Silver nitrate (AgNO<sub>3</sub>) is used to make several different silver halides used in photographic films. How many formula units of AgNO<sub>3</sub> are there in 3.25 mol AgNO<sub>3</sub>?

$$3.25 \text{ mol AgNO}_3 \times \frac{6.02 \times 10^{23} \text{ formula units AgNO}_3}{1 \text{ mol AgNO}_3} \\ = 1.96 \times 10^{24} \text{ formula units AgNO}_3$$

4. **Challenge** Calculate the number of oxygen atoms in 5.00 mol of oxygen molecules. Oxygen is a diatomic molecule, O<sub>2</sub>.

$$5.00 \text{ mol O}_2 \times \frac{6.02 \times 10^{23} \text{ molecules O}_2}{1 \text{ mol O}_2} \\ \times \frac{2 \text{ O atoms}}{\text{molecule O}_2} = 6.02 \times 10^{24} \text{ atoms O}$$

5. How many moles contain each of the following?
  - a.  $5.75 \times 10^{24}$  atoms Al

$$5.75 \times 10^{24} \text{ atoms Al} \times \frac{1 \text{ mol Al}}{6.02 \times 10^{23} \text{ atoms Al}} \\ = 9.55 \text{ mol Al}$$

- b.  $2.50 \times 10^{20}$  atoms Fe

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$$2.50 \times 10^{20} \text{ atoms Fe} \times \frac{1 \text{ mol Fe}}{6.02 \times 10^{23} \text{ atoms Fe}} \\ = 4.15 \times 10^{-4} \text{ mol Fe}$$

6. **Challenge** Identify the representative particle for each formula and convert the given number of representative particles to moles.

- a.  $3.75 \times 10^{24}$  CO<sub>2</sub>

The representative particle is a molecule.

$$3.75 \times 10^{24} \text{ molecules CO}_2 \times \frac{1 \text{ mol CO}_2}{6.02 \times 10^{23} \text{ molecules CO}_2} = 6.23 \text{ mol CO}_2$$

- b.  $3.58 \times 10^{23}$  ZnCl<sub>2</sub>

The representative particle is a formula unit.

$$3.58 \times 10^{23} \text{ formula units ZnCl}_2 \times \frac{1 \text{ mol ZnCl}_2}{6.02 \times 10^{23} \text{ formula units ZnCl}_2} \\ = 0.595 \text{ mol ZnCl}_2$$

## Section 1 Review

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7. **Explain** why chemists use the mole.

Chemists use the mole because it is a convenient way of knowing how many representative particles are in a sample.

8. **State** the mathematical relationship between Avogadro's number and 1 mol.

One mole contains Avogadro's number ( $6.02 \times 10^{23}$ ) of representative particles.

# Mole Solution Manual

**C. A. Trapp, Peter Atkins, Julio dePaula**



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chemical reaction of phosphorus with oxygen is partly responsible for the spectacular show of light Carbon The element is found in several forms including the familiar diamond and another recently discovered sooty substance that consists of soccer ball shaped molecules often referred to as buckeyballs Diamond is not the most stable form of carbon and is created from other forms of carbon at high temperatures and pressures deep within the earth Acetylene torch Cutting steel is possible because of the intense heat generated by the chemical reaction of acetylene with oxygen a reaction between molecules of  $C_2H_2$  and  $O_2$  to give  $CO_2$  and  $H_2O$  Hot air balloon The air that helps it rise is heated by the combustion of molecules of propane each composed of three carbon and eight hydrogen atoms Stormy weather The evaporation of water serves to store energy provided by the sun Subsequent condensation of the water vapor releases this energy and is the basis of all the weather systems on our planet

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