

Chemistry 12

Solutions Manual Part B

Unit 3 Energy Changes and Rates of Reaction

Solutions to Practice Problems in Chapter 5 Energy Changes

Calculating the Absorption of Heat (Student textbook page 281)

1. How much heat must be added to a 36.2 g sample of iron to increase its temperature by 250.0°C?

What Is Required?

You have to calculate the amount of heat that must be added to a sample of iron to increase its temperature by 250.0°C.

What Is Given?

You know the mass of iron: $m = 36.2 \text{ g}$

You know the change in the temperature: $\Delta T = 250.0^\circ\text{C}$

You have the specific heat capacity of iron from **Table 5.1** on page 280 of the student text: $c_{\text{Fe}} = 0.449 \text{ J/g}^\circ\text{C}$

Plan Your Strategy	Act on Your Strategy
Substitute the given information into the formula $Q = mc\Delta T$. Solve for Q .	$\begin{aligned} Q &= mc\Delta T \\ &= (36.2 \text{ g}) \left(0.449 \frac{\text{J}}{\text{g}^\circ\text{C}} \right) (250.0^\circ\text{C}) \\ &= 4063.45 \text{ J} \\ &= 4.06 \text{ kJ} \end{aligned}$ <p>The iron must gain 4060 J or 4.06 kJ of heat energy.</p>

Check Your Solution

Check that the information has been substituted correctly. The answer has the correct number of significant digits.

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Table of Contents Mhr Chemistry 12 Solutions

1. Understanding the eBook Mhr Chemistry 12 Solutions
 - The Rise of Digital Reading Mhr Chemistry 12 Solutions
 - Advantages of eBooks Over Traditional Books
2. Identifying Mhr Chemistry 12 Solutions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mhr Chemistry 12 Solutions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mhr Chemistry 12 Solutions
 - Personalized Recommendations
 - Mhr Chemistry 12 Solutions User Reviews and Ratings
 - Mhr Chemistry 12 Solutions and Bestseller Lists
5. Accessing Mhr Chemistry 12 Solutions Free and Paid eBooks
 - Mhr Chemistry 12 Solutions Public Domain eBooks
 - Mhr Chemistry 12 Solutions eBook Subscription Services
 - Mhr Chemistry 12 Solutions Budget-Friendly Options
6. Navigating Mhr Chemistry 12 Solutions eBook Formats

- ePub, PDF, MOBI, and More
- Mhr Chemistry 12 Solutions Compatibility with Devices
- Mhr Chemistry 12 Solutions Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mhr Chemistry 12 Solutions
 - Highlighting and Note-Taking Mhr Chemistry 12 Solutions
 - Interactive Elements Mhr Chemistry 12 Solutions
- 8. Staying Engaged with Mhr Chemistry 12 Solutions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mhr Chemistry 12 Solutions
- 9. Balancing eBooks and Physical Books Mhr Chemistry 12 Solutions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mhr Chemistry 12 Solutions
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Mhr Chemistry 12 Solutions
 - Setting Reading Goals Mhr Chemistry 12 Solutions
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mhr Chemistry 12 Solutions
 - Fact-Checking eBook Content of Mhr Chemistry 12 Solutions
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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