



MOTOROLA Repeater Interface

REPEATER  
ENABLE

CCR

SET-UP

REPEATER RX

CDM1550-LS\*

P1



P3

P2



P4



MOTOROLA

REPEATER TX

CDM1550-LS\*

P1



P3

P2



P4



MOTOROLA



MOTOROLA

CDR700

# Motorola Cdr700 Repeater Manual

**Huangqi Zhang**



**Motorola Cdr700 Repeater Manual:**

**MTR300 Mountain Top Repeater** Motorola Canada Limited,1985      **Motorola Service Manual** Motorola, Inc,195?  
Motorola RF Data Manual Motorola Semiconductor Products Inc. Technical Information Center,1978      **Motorola RF**  
**Data Manual** Motorola, Inc,1980      *Operator, Technical, Installation and Service Manual - Universal Digital Repeater*  
,1996      **Specialized Auto Radio Manual, 6-A** John Francis Rider,1955      **Motorola Inc. 25-44 Megacycles 50 Watt**  
**Fixed Stations Instruction Manual** ,1948      Instruction Manual Motorola the "research" Line of Radio Communication  
Systems ,1955      *Philco-Delco Car Radio Installation and Service Manual* PHILCO-DELCO CAR RADIO INSTALLATION AND  
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web elements in the second row of the periodic table place their electrons in the 2n shell as well as the 1n shell for instance lithium li has three electrons two fill the 1 s orbital and the third is placed in the 2 s orbital giving an electron configuration of 1 s<sup>2</sup> 2 s<sup>1</sup>

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web electron configurations and trends answer these discussion questions given the following part of an electron

configuration 3s<sup>2</sup> explain which part refers to the number of electrons the energy level and the sublevel what is the maximum number of electrons that can occupy a single orbital state and explain hund's rule

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web one more point needs to be emphasized about the relationship between electron configuration and the periodic table the atoms of elements in the same vertical column of the table have similar electron configurations for example consider the alkaline earth elements group iia

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web the electron removed during the ionization of beryllium he 2s<sup>2</sup> is an s electron whereas the electron removed during the ionization of boron he 2s<sup>2</sup> 2p<sup>1</sup> is a p electron this results in a lower first ionization energy for boron even though its nuclear charge is greater by one proton

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web predictions from reliable sources have been used for these elements grayed out electron numbers indicate subshells filled to their maximum bracketed noble gas symbols on the left represent inner configurations that are the same in each period written out these are he 2 helium 1s<sup>2</sup> ne 10 neon 1s<sup>2</sup> 2s<sup>2</sup> 2p<sup>6</sup>

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