

15. The equilibrium constant for the reaction:



at 448°C is 50.

a) If 1.0 mol of H<sub>2</sub> is mixed with 1.0 mol of I<sub>2</sub> in a 0.50 L container and allowed to react at 448°C, what is the equilibrium [HI]?

Initial [H<sub>2</sub>] =  $\frac{1.0 \text{ mol}}{0.50 \text{ L}} = 2.0 \text{ M}$ , [I<sub>2</sub>] = 2.0 M, [HI] = 0

	$\text{H}_2 + \text{I}_2 \rightleftharpoons 2\text{HI}$		
[I]	2.0	2.0	0
[C]	-x	-x	+2x
[E]	2.0-x	2.0-x	2x
[E]	2.0-1.559	2.0-1.559	2(1.559)
[E]	0.44	0.44	3.1

$$K_{eq} = \frac{[\text{HI}]^2}{[\text{H}_2][\text{I}_2]}$$

$$50. = \frac{(2x)^2}{(2.0-x)^2}$$

$$\sqrt{50.} = \frac{2x}{2-x}$$

$$\begin{aligned} 7.071(2-x) &= 2x \\ 14.142 - 7.071x &= 2x \\ 14.142 &= 9.071x \\ x &= \frac{14.142}{9.071} = 1.559 \end{aligned}$$

Answer [HI] = 3.1 M

b) How many moles of HI are formed at equilibrium? (Actual yield)

$$3.118 \text{ M} \times 0.50 \text{ L} = 1.6 \text{ mol}$$

Answer 1.6 moles of HI

16. Given K<sub>eq</sub> for the reaction:



is 0.042 at 250°C, what will happen if 2.50 mol of PCl<sub>5</sub>, 0.600 mol of Cl<sub>2</sub> and 0.600 mol of PCl<sub>3</sub> are placed in a 1.00 flask at 250°C? (If all the reaction shift left, right, or not occur at all?)

$$K_{eq} = \frac{[\text{PCl}_3][\text{Cl}_2]}{[\text{PCl}_5]}$$

$$\text{Trial } K_{eq} = \frac{(0.600)(0.600)}{(2.50)}$$

$$\text{Trial } K_{eq} = 0.144$$

$$\text{Actual } K_{eq} = 0.042$$

Since Trial K<sub>eq</sub> > K<sub>eq</sub>  
The reaction will shift to the LEFT as it approaches equilibrium.  
So [PCl<sub>5</sub>] will increase while [PCl<sub>3</sub>] and [Cl<sub>2</sub>] decrease.

Answer shift LEFT

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## **Table of Contents Nelson Chemistry Exercise Answers**

1. Understanding the eBook Nelson Chemistry Exercise Answers
  - The Rise of Digital Reading Nelson Chemistry Exercise Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Nelson Chemistry Exercise Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an eBook Platform
  - User-Friendly Interface
4. Exploring eBook Recommendations from Nelson Chemistry Exercise Answers
  - Personalized Recommendations
  - Nelson Chemistry Exercise Answers User Reviews and Ratings
  - Nelson Chemistry Exercise Answers and Bestseller Lists
5. Accessing Nelson Chemistry Exercise Answers Free and Paid eBooks
  - Nelson Chemistry Exercise Answers Public Domain eBooks
  - Nelson Chemistry Exercise Answers eBook Subscription Services
  - Nelson Chemistry Exercise Answers Budget-Friendly Options
6. Navigating Nelson Chemistry Exercise Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Nelson Chemistry Exercise Answers Compatibility with Devices
  - Nelson Chemistry Exercise Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Nelson Chemistry Exercise Answers
  - Highlighting and Note-Taking Nelson Chemistry Exercise Answers
  - Interactive Elements Nelson Chemistry Exercise Answers
8. Staying Engaged with Nelson Chemistry Exercise Answers

- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Nelson Chemistry Exercise Answers
9. Balancing eBooks and Physical Books Nelson Chemistry Exercise Answers
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Nelson Chemistry Exercise Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Nelson Chemistry Exercise Answers
- Setting Reading Goals Nelson Chemistry Exercise Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Nelson Chemistry Exercise Answers
- Fact-Checking eBook Content of Nelson Chemistry Exercise Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
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