

Acids, Bases, and Salts

Arrhenius Theory

Base: A substance that produces hydroxide ions (OH^-) in aqueous solution.

Table 7.2 Some Common Bases

Name	Formula	Classification	Common Uses/Notes
Sodium hydroxide	NaOH	Strong	Acid neutralization; soap making
Potassium hydroxide	KOH	Strong	Making liquid soaps and biodiesel fuels
Lithium hydroxide	LiOH	Strong	Alkaline storage batteries
Calcium hydroxide	$\text{Ca}(\text{OH})_2$	Strong ^a	Plaster, cement; soil neutralizer
Magnesium hydroxide	$\text{Mg}(\text{OH})_2$	Strong ^a	Antacid, laxative
Ammonia	NH_3	Weak	Fertilizer, household cleansers

^aAlthough these bases are classified as strong, they are not very soluble in water. Calcium hydroxide is only slightly soluble, and magnesium hydroxide is practically insoluble.

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