

## Chemistry CH 13 solids and modern materials

- Graphene- new modern solid compound.
  - Carbon solid
  - Recently discovered
  - Thin as thin as one atom sheet
  - Very strong (strongest known material)
  - Resistant to the heat
  - Still trying to find applications for it like in computers, plastics and satellites
- X-ray crystallography- technique used to **find structure of any solids**
  - If we pass the x-ray light through the crystal it shows bond angles, size of atoms
  - Crystal lattice
    - Unit cell- smallest atom, repeating molecules/atom/particle within crystal.
      - Different types of unit cells
      - In cubic= all bond lengths and all bond angles are equal
        - Simple cubic: 1/8 of each corner is of atom
          - There is one atom present in
          - $L = 2\text{radius}$
          - Coordination number=6
          - Packing efficiency = 52%
        - Body-centered cubic
          - 1/8 of each corner particle lies in cell
          - And there is one atom inside the cube too
          - Packing efficiency= 68%
          - Coordination number = 8
          - $L=4r$
        - Face-centered cubic
          - Inside the particle there are 3 atoms inside
          - And 1/8 of each corner so total= 4 atoms pithing
          - Packing efficiency = 74%
          - Coordination number = 12
      - Coordination number- number of atoms or molecules surrounding it or the atoms that are touching the central atom.
      - Packing efficiency- depends on coordination number, percentage of volume unit cells occupy.
      - 
      - Classification of Crystalline solids
        - Molecular solids- composed of molecules
          - some crystals are called polymorphs- having different properties
        - Ionic solids- composed of ions
          - Anions are larger than cations because of the negative charge
          - Cesium chloride- body centered cs in the center and Cl Surrounds it

# Pearson Hall Chemistry Ch 13 Guided

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