

# SOLUTIONS MANUAL

## OPERATING SYSTEMS: INTERNALS AND DESIGN PRINCIPLES FIFTH EDITION

WILLIAM STALLINGS

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# Operating System Solution Manual

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$$4\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Cu} + 2\text{Ag} + 2\text{Cu(NO}_3)_2$$

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calculated mass of the product expected  $\times 100$  in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction  $\text{Al}_2\text{O}_3 + 4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$  b determine the theoretical yield of  $\text{Al}_2\text{O}_3$  160.0 g  $\text{Al}$   $\times \frac{1 \text{ mol Al}}{27 \text{ g Al}} \times 2 \text{ mol Al}_2\text{O}_3 \times 101.964 \text{ g Al}_2\text{O}_3 = 302.3 \text{ g Al}_2\text{O}_3$  c determine the percent yield  $\frac{\text{actual}}{\text{theoretical}} \times 100 = \frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$

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web 7  $\text{Na}_2\text{CO}_4$  is the limiting reactant percent yield 86.68 only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12  $\text{C}$  is the limiting reagent 4.33 g of  $\text{H}_2$  are left over 13

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web the percent yield is calculated as follows  $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$  percent yield theoretical  $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$  based on this definition we would expect a percent yield

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which reactant is limiting assuming we started with 30.0 grams of ammonium nitrate and 50.0 grams of sodium phosphate what is the

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80 g  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2$  xs 1 333.8 g  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2\text{O}_5$  28 g  $\text{CO}$  1 mol  $\text{CO}$

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