

Chemistry Chapter 14a - Thermodynamics

- Be able to predict the spontaneity of a reaction from either its standard Gibbs free energy or its standard enthalpy and entropy values (don't forget to look at the units!)
- Remember that following definitions are important for exams:

Standard Gibbs Free Energy for Reaction

Use the following standard thermodynamic properties of substances to calculate ΔG° for a reaction and then calculate ΔG for a reaction under non-standard conditions (remember that ΔG is the Gibbs free energy of a reaction and ΔG° is the standard Gibbs free energy of a reaction).

Substance	ΔG°_f (kJ/mol)
$\text{H}_2\text{O}(l)$	-237.1
$\text{H}_2\text{O}(g)$	-228.6
$\text{H}_2(g)$	0
$\text{O}_2(g)$	0
$\text{CO}_2(g)$	-394.4
$\text{CH}_4(g)$	-50.7
$\text{C}_2\text{H}_6(g)$	-67.1
$\text{C}_2\text{H}_4(g)$	68.1
$\text{C}_2\text{H}_2(g)$	209.2
$\text{C}_2\text{H}_6(g)$	-67.1
$\text{C}_2\text{H}_4(g)$	68.1
$\text{C}_2\text{H}_2(g)$	209.2
$\text{C}_2\text{H}_6(g)$	-67.1
$\text{C}_2\text{H}_4(g)$	68.1
$\text{C}_2\text{H}_2(g)$	209.2

Substance	ΔH°_f (kJ/mol)
$\text{H}_2\text{O}(l)$	-285.8
$\text{H}_2\text{O}(g)$	-241.8
$\text{H}_2(g)$	0
$\text{O}_2(g)$	0
$\text{CO}_2(g)$	-393.5
$\text{CH}_4(g)$	-74.6
$\text{C}_2\text{H}_6(g)$	-84.7
$\text{C}_2\text{H}_4(g)$	52.3
$\text{C}_2\text{H}_2(g)$	226.7
$\text{C}_2\text{H}_6(g)$	-84.7
$\text{C}_2\text{H}_4(g)$	52.3
$\text{C}_2\text{H}_2(g)$	226.7

Step 1: Calculate ΔG°

Step 2: Calculate ΔG

Use the following information to get ΔG° for a reaction and then calculate ΔG for a reaction under non-standard conditions.

- ΔG°_f for $\text{H}_2\text{O}(l)$ is -237.1 kJ/mol
- ΔG°_f for $\text{H}_2\text{O}(g)$ is -228.6 kJ/mol
- ΔG°_f for $\text{H}_2(g)$ is 0 kJ/mol
- ΔG°_f for $\text{O}_2(g)$ is 0 kJ/mol
- ΔG°_f for $\text{CO}_2(g)$ is -394.4 kJ/mol
- ΔG°_f for $\text{CH}_4(g)$ is -50.7 kJ/mol
- ΔG°_f for $\text{C}_2\text{H}_6(g)$ is -67.1 kJ/mol
- ΔG°_f for $\text{C}_2\text{H}_4(g)$ is 68.1 kJ/mol
- ΔG°_f for $\text{C}_2\text{H}_2(g)$ is 209.2 kJ/mol
- ΔG°_f for $\text{C}_2\text{H}_6(g)$ is -67.1 kJ/mol
- ΔG°_f for $\text{C}_2\text{H}_4(g)$ is 68.1 kJ/mol
- ΔG°_f for $\text{C}_2\text{H}_2(g)$ is 209.2 kJ/mol

- ΔH°_f for $\text{H}_2\text{O}(l)$ is -285.8 kJ/mol
- ΔH°_f for $\text{H}_2\text{O}(g)$ is -241.8 kJ/mol
- ΔH°_f for $\text{H}_2(g)$ is 0 kJ/mol
- ΔH°_f for $\text{O}_2(g)$ is 0 kJ/mol
- ΔH°_f for $\text{CO}_2(g)$ is -393.5 kJ/mol
- ΔH°_f for $\text{CH}_4(g)$ is -74.6 kJ/mol
- ΔH°_f for $\text{C}_2\text{H}_6(g)$ is -84.7 kJ/mol
- ΔH°_f for $\text{C}_2\text{H}_4(g)$ is 52.3 kJ/mol
- ΔH°_f for $\text{C}_2\text{H}_2(g)$ is 226.7 kJ/mol
- ΔH°_f for $\text{C}_2\text{H}_6(g)$ is -84.7 kJ/mol
- ΔH°_f for $\text{C}_2\text{H}_4(g)$ is 52.3 kJ/mol
- ΔH°_f for $\text{C}_2\text{H}_2(g)$ is 226.7 kJ/mol

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