

Reactions of Amines and Epoxides

Preparation of Amines

Amination of alcohols: Polyalkylation is a serious problem for indirect synthesis of amines



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Bond Types, Strength, Length, Polarity and Dipole Moment

Two types of covalent bond

There are two types of covalent bond: bonding and non-bonding bond

- Bonding bonds are formed by the overlap of electron orbitals and are classified as sigma (stronger) and pi (weaker) bonds.



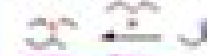
- Non-bonding bonds are formed by the overlap of electron orbitals between two atoms.



The difference between sigma and pi bonds



Bonding and Non-bonding bonds



Bonding energy



There are two types of covalent bond



Sigma and pi bonds

Covalent bonds are classified as sigma and pi bonds

- Sigma bonds are formed by the end-to-end overlap of orbitals.
- Pi bonds are formed by the side-to-side overlap of orbitals.
- Sigma bonds are stronger than pi bonds.
- Pi bonds are weaker than sigma bonds.



Pi bonds are weaker than sigma bonds

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds can be polar and nonpolar.



- If the difference in electronegativity is less than 0.5, the electrons are shared equally (nonpolar bond).
- If the difference in electronegativity is between 0.5 and 1.7, the electrons are shared unequally (polar bond).
- If the difference in electronegativity is greater than 1.7, the electrons are not shared (ionic bond).

- A difference of 0.5 or higher can be high for the electrons are not shared (ionic bond).

- A difference of 1.7 or higher can be high for the electrons are not shared (ionic bond).

Dipole Moment

The polarity of a bond is described by the dipole moment (μ)



- The dipole moment is a vector pointing from the less electronegative atom to the more electronegative atom.
- The vector dipole moment is a vector pointing from the less electronegative atom to the more electronegative atom.
- The dipole moment is a vector pointing from the less electronegative atom to the more electronegative atom.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipole moments

The polarity of the molecule is determined by the vector sum of the dipole moments. A polar molecule has a net dipole moment.



Vector sum of the dipole moments is the molecular dipole moment

Bond Length and Bond Strength

The bond length is the distance between the nuclei of the two atoms.

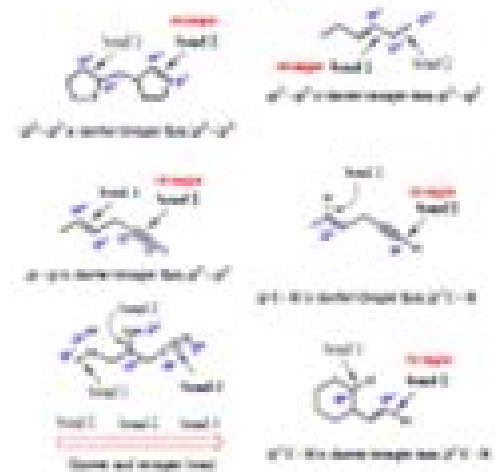


The bond length is the distance between the nuclei of the two atoms.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	1.54 Å	1.34 Å	1.20 Å
Strength	33 kcal/mol	47 kcal/mol	58 kcal/mol
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
C-H Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	1.09 Å	1.08 Å	1.06 Å
Strength	99 kcal/mol	112 kcal/mol	120 kcal/mol

Higher s character = stronger and shorter bond

Examples



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