

Reactions of Amines and Epoxides

Preparation of Amines

Amination of alcohols: Polyalkylation is a serious problem for indirect synthesis of amines



A Summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Bond Types, Strength, Length, Polarity and Dipole Moment

Two types of covalent bond

There are two types of covalent bond: bonding and non-bonding bond

• **Bonding bonds** are formed by the overlap of atomic orbitals from one element to another (homonuclear) (positive) and another (heteronuclear) (negative) (negative)



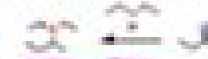
• **Non-bonding bonds** are formed by the overlap of atomic orbitals between two atoms



The Relative Electronegativity of Atoms



Bonding and Non-bonding Bonds



Bonding and Non-bonding Bonds



There are two types of covalent bond

Definition of covalent bond



Sigma and pi bonds

Covalent bonds are classified in sigma and pi bonds

• **Sigma bonds** are formed by the overlap of atomic orbitals

• **pi bonds** are formed by the overlap of atomic orbitals

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Dipole Moment

The polarity of a bond is described by the dipole moment (μ)



• The dipole moment is defined as a vector pointing from the less electronegative atom to the more electronegative atom

• The vector dipole moment is also a vector pointing from the less electronegative atom to the more electronegative atom

• μ (dipole moment) indicates an electron density shift

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Molecular Dipole

To determine the molecular polarity, the vector sum of the dipole moment

The polarity of the molecule is determined by the vector sum of the dipole moment

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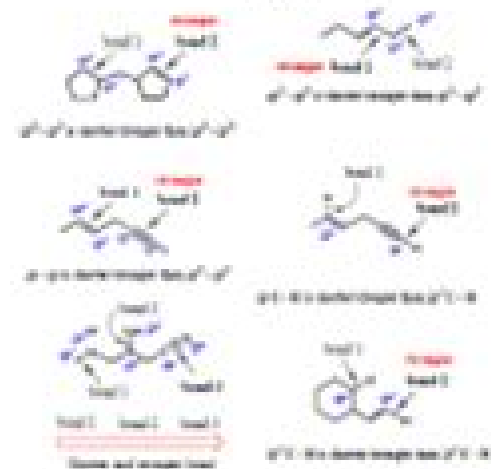
The polarity of the molecule is determined by the vector sum of the dipole moment

The polarity of the molecule is determined by the vector sum of the dipole moment

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	1.54 Å	1.34 Å	1.20 Å
Strength	33 kcal/mol	112 kcal/mol	200 kcal/mol
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
C-H Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	1.09 Å	1.08 Å	1.06 Å
Strength	98 kcal/mol	110 kcal/mol	120 kcal/mol

Higher s character = stronger and shorter bond

Examples



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The Princeton Review



Organic Chemistry Study Guide For Pcat:

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- Chapter 3: Organic Chemistry Study Guide For Pcat in Everyday Life
- Chapter 4: Organic Chemistry Study Guide For Pcat in Specific Contexts
- Chapter 5: Conclusion

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