

C. Melting Point Determination

Recrystallized Benzoic acid	Tube	Melting Point (°C)	Complete Liquidation (°C)	Melting point range (°C)	Change in color	Shrinking of size
	1	122	127	122-127	Simultaneously white to colorless	Gradually shrinks up to melting point
	2	122	127	122-127	Simultaneously white to colorless	Gradually shrinks up to melting point

Benzoic samples	Melting point range (°C)
Recrystallized	122-127
Crude	126-128

- As the crystals were heated, there is a continuous but gradual shrinking of the size of the crystals as evidenced by the diminishing or decreasing height of the sample observed in the capillary tubes for both trials accompanied by a change in color as indicated. Also, it was observed that after removal of the capillary tubes after each trial from the hot oil bath, the purified benzoic crystals immediately solidified, that is, returned to the original solid phase.

IV. Discussion and Interpretation of Results

Recrystallization from a solvent depends primarily on the fact that more solids are more soluble in hot solvents than cold solvents, and second, different materials are soluble in different solvents in various solvents. In order for recrystallization to be successful, there should be a large difference in the solubility of the solid material in hot solvents compared to cold solvents. A material can be purified by recrystallization when both the desired substance and the impurity have similar solubilities, but only when the impurity represents a small fraction of the total solid. The desired substance will crystallize on cooling, but the impurity will not.

The most crucial and time-consuming feature of recrystallization is the selection of solvent because important considerations should be observed. In order to choose a good recrystallization solvent, the solubility of the solute should be compared at its other words, sparingly soluble only in the solvent at room temperature. Thus, from this relative description, we have to consider the principle "like dissolves like" which refers to the polarity of both the solute and the solvent. Substances of same polarity tend to dissolve one another more at room temperature, which therefore contributes the very first condition of choosing a good recrystallization solvent.

In the experiment, benzoic acid, whose structural formula is given below, was used as the sample.

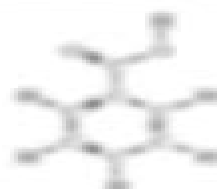


Figure 1. Structural formula of Benzoic acid

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John R. Holum, Ruth C. Denison



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